

Chemistry 1330

Personal Performance Assessment #3

Professor Keller

Name

KEY

Student #

Lab Section

Multiple Choice (4 points each). Circle the one best answer for the following questions.

- The pH of neutral water is 6.4 at 80 °C. From this we can infer that the auto-dissociation of water is _____ and that K_w _____ with increasing temperature.

(A) endothermic, increases (B) endothermic, decreases
(C) exothermic, increases (D) exothermic, decrease
- Which cannot act as a Brønsted-Lowry base?

(A) H₂O (B) CH₄ (C) NH₃ (D) OCl⁻ (E) none of them

No lone pair
- If $K_{eq} \gg 1$ for a reaction, then

(A) $\Delta G^\circ \ll 1$ (B) $0 < \Delta G^\circ < 1$ (C) $\Delta G^\circ < 0$ (D) $\Delta G^\circ \gg 1$ (E) $\Delta G^\circ = 0$

(C) is best, but either OK.
- Which statement is correct?

(A) The energy of the Universe is always increasing.
(B) Processes having $\Delta G^\circ < 0$ always result in an increase in entropy of the system.
(C) Endothermic processes decrease the entropy of the surroundings
(D) Exothermic reactions always have $\Delta G^\circ < 0$.
(E) None of the above statements are correct.
- Consider the product-favored reaction: $S_2(g) + 2 O_2(g) \rightleftharpoons 2 SO_2(g)$. What are the signs of ΔH° , ΔS° , and ΔG° for this reaction?

(A) $\Delta H^\circ < 0$; $\Delta S^\circ < 0$; $\Delta G^\circ < 0$ (B) $\Delta H^\circ < 0$; $\Delta S^\circ < 0$; $\Delta G^\circ > 0$
(C) $\Delta H^\circ > 0$; $\Delta S^\circ < 0$; $\Delta G^\circ > 0$ (D) $\Delta H^\circ > 0$; $\Delta S^\circ > 0$; $\Delta G^\circ < 0$
- The entropy of the system will increase when
 - a molecule is broken into two or more smaller molecules.
 - a reaction occurs that results in an increase in the number of moles of gas.
 - a solid changes to a liquid.
 - a gas changes to a liquid.

(A) I and III only (b) II and IV only (c) I and II only (D) I, II, and III only

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Scott C. Dulebohn



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