

## Lab 22A

### Investigating Chemical Equilibrium

**Purpose:** To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a function of concentration and temperature and make, via the system, visible use of Le Chatelier's principle through applying Le Chatelier's principle.

#### Proble:

1. Write a balanced equation for the reaction in part 1.  
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
2. Write a balanced equation for the reaction in part 2.  
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
3. Write a balanced equation for the reaction in part 3.  
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
4. In part 1, what is the purpose of adding potassium thiocyanate to the solution of iron(III) ions?  
To observe shifts in the equilibrium and to use Le Chatelier's principle to predict the direction of the shift. The purpose of adding iron(III) ions to the solution of thiocyanate ions is to observe shifts in the equilibrium and to use Le Chatelier's principle to predict the direction of the shift.
5. What is the purpose of adding distilled water to part 1?  
To observe shifts in the equilibrium and to use Le Chatelier's principle to predict the direction of the shift.
6. What is the purpose of adding distilled water to part 2?  
To observe shifts in the equilibrium and to use Le Chatelier's principle to predict the direction of the shift.
7. Write the equation for the reaction between  $\text{Fe}^{3+}$  and  $\text{OH}^{-}$  ions.  
$$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})_3(\text{s})$$
8. Write the equation for the reaction between  $\text{Fe}^{3+}$  and  $\text{OH}^{-}$  ions.  
$$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})_3(\text{s})$$
9. What is the purpose of adding distilled water to part 7?  
To observe shifts in the equilibrium and to use Le Chatelier's principle to predict the direction of the shift.

# Lab 12a Investigating Chemical Equilibrium

**RS Peters**



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