

# ACID-BASE RELATIVE STRENGTHS in Aqueous Solutions at 25° C

Name of Acid	Acid		Base	$K_a$
Hydriodic	HI	→	H <sup>+</sup> + I <sup>-</sup>	≫ 1
Hydrobromic	HBr	→	H <sup>+</sup> + Br <sup>-</sup>	≫ 1
Hydrochloric	HCl	→	H <sup>+</sup> + Cl <sup>-</sup>	≫ 1
Nitric	HNO <sub>3</sub>	→	H <sup>+</sup> + NO <sub>3</sub> <sup>-</sup>	> 1
Sulfuric	H <sub>2</sub> SO <sub>4</sub>	→	H <sup>+</sup> + HSO <sub>4</sub> <sup>-</sup>	> 1
Oxalic	H <sub>2</sub> C <sub>2</sub> O <sub>4</sub>	⇌	H <sup>+</sup> + HC <sub>2</sub> O <sub>4</sub> <sup>-</sup>	5.90 × 10 <sup>-2</sup>
Sulfurous	H <sub>2</sub> SO <sub>3</sub>	⇌	H <sup>+</sup> + HSO <sub>3</sub> <sup>-</sup>	1.54 × 10 <sup>-2</sup>
Hydrogen Sulfate Ion	HSO <sub>4</sub> <sup>-</sup>	⇌	H <sup>+</sup> + SO <sub>4</sub> <sup>2-</sup>	1.20 × 10 <sup>-2</sup>
Phosphoric	H <sub>3</sub> PO <sub>4</sub>	⇌	H <sup>+</sup> + H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>	7.52 × 10 <sup>-3</sup>
Nitrous	HNO <sub>2</sub>	⇌	H <sup>+</sup> + NO <sub>2</sub> <sup>-</sup>	4.6 × 10 <sup>-4</sup>
Hydrofluoric	HF	⇌	H <sup>+</sup> + F <sup>-</sup>	3.53 × 10 <sup>-4</sup>
Formic	HCO <sub>2</sub> H	⇌	H <sup>+</sup> + HCO <sub>2</sub> <sup>-</sup>	1.77 × 10 <sup>-4</sup>
Benzoic	C <sub>6</sub> H <sub>5</sub> CO <sub>2</sub> H	⇌	H <sup>+</sup> + C <sub>6</sub> H <sub>5</sub> CO <sub>2</sub> <sup>-</sup>	6.46 × 10 <sup>-5</sup>
Acetic	CH <sub>3</sub> CO <sub>2</sub> H	⇌	H <sup>+</sup> + CH <sub>3</sub> CO <sub>2</sub> <sup>-</sup>	1.76 × 10 <sup>-5</sup>
Carbonic	H <sub>2</sub> CO <sub>3</sub>	⇌	H <sup>+</sup> + HCO <sub>3</sub> <sup>-</sup>	4.30 × 10 <sup>-7</sup>
Hydrogen Sulfite Ion	HSO <sub>3</sub> <sup>-</sup>	⇌	H <sup>+</sup> + SO <sub>3</sub> <sup>2-</sup>	1.02 × 10 <sup>-7</sup>
Dihydrogen Phosphate Ion	H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>	⇌	H <sup>+</sup> + HPO <sub>4</sub> <sup>2-</sup>	6.23 × 10 <sup>-8</sup>
Boric	H <sub>3</sub> BO <sub>3</sub>	⇌	H <sup>+</sup> + H <sub>2</sub> BO <sub>3</sub> <sup>-</sup>	7.3 × 10 <sup>-10</sup>
Ammonium Ion	NH <sub>4</sub> <sup>+</sup>	⇌	H <sup>+</sup> + NH <sub>4</sub> <sup>+</sup>	5.64 × 10 <sup>-10</sup>
Hydrogen Carbonate Ion	HCO <sub>3</sub> <sup>-</sup>	⇌	H <sup>+</sup> + CO <sub>3</sub> <sup>2-</sup>	5.61 × 10 <sup>-11</sup>
Monohydrogen Phosphate Ion	HPO <sub>4</sub> <sup>2-</sup>	⇌	H <sup>+</sup> + PO <sub>4</sub> <sup>3-</sup>	2.2 × 10 <sup>-13</sup>
Water	H <sub>2</sub> O	⇌	H <sup>+</sup> + OH <sup>-</sup>	1.0 × 10 <sup>-14</sup>

Strongest Acid

Weak Acid

Weakest Acid

Weakest Base

Weak Base

Strongest Base

For space considerations, all reactions omit water as a reactant  
 $HA + H_2O \rightleftharpoons H_3O^+ + A^-$

# Lab 16 1 Creating Acid Base Strength Table

**M Planty**



**Lab 16 1 Creating Acid Base Strength Table:**

## Reviewing **Lab 16 1 Creating Acid Base Strength Table**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is really astonishing. Within the pages of "**Lab 16 1 Creating Acid Base Strength Table**," an enthralling opus penned by a very acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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