

Practical Problems Chapter 9: Solutions

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2. The following procedure illustrates how to determine the molar mass of a volatile liquid. A student used the following procedure to determine the molar mass of a substance with a density of 1.10 g/mL. What is the molar mass of the substance? (M_w = 100 g/mol)

1. Weighing flask

$$\text{Initial mass (g)} = \left(\frac{10.0 \text{ g water}}{1 \text{ mL water}} \right) \times 10.0 \text{ mL water}$$

$$10.0 \left(\frac{10.0}{10.0} \right) = 10.0 \left(\frac{10.0}{10.0} \right) = \frac{100.0 \text{ g water}}{10.0 \text{ mL water}} \times 10.0 \text{ mL water} = 100.0 \text{ g water} \left(\frac{10.0}{10.0} \right) \text{ mL water}$$

2. Mass flask

$$10.0 \left(\frac{10.0}{10.0} \right) = 10.0 \left(\frac{10.0}{10.0} \right) = \frac{100.0 \text{ g water}}{10.0 \text{ mL water}} \times 10.0 \text{ mL water} = 100.0 \text{ g water} \left(\frac{10.0}{10.0} \right) \text{ mL water}$$

3. Mass flask

$$\text{Molar mass} = \left(\frac{100.0 \text{ g water}}{10.0 \text{ mL water}} \right) \times 10.0 \text{ mL water}$$

$$\text{Molar mass} = \left(\frac{100.0 \text{ g water}}{10.0 \text{ mL water}} \right) \times 10.0 \text{ mL water}$$

$$\text{Molar mass} = 10.0 \text{ mL water} \times \frac{100.0 \text{ g water}}{10.0 \text{ mL water}} = 100.0 \text{ g water}$$

10. (points) 10.11 to 10.14 (10.11)

11. (points) 10.15 to 10.18 (10.15)

12. (points) 10.19 to 10.22 (10.19)

23. How many grams of Na_2CO_3 are represented under the flask on a 100 mL beaker of Na_2CO_3 ?

$$\text{Initial mass (g)} = \left(\frac{10.0 \text{ g water}}{1 \text{ mL water}} \right) \times \left(\frac{10.0 \text{ mL water}}{1 \text{ mL water}} \right) \times \left(\frac{10.0 \text{ mL water}}{1 \text{ mL water}} \right) = 100.0 \text{ g water}$$

Answer: 100.0 g Na_2CO_3

Guided Practice Problems Chapter 8 Chemistry

LM Reder

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Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

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