

CHAPTER 6

STUDY GUIDE

Ionic Compounds and Metals

Lesson 6.1 Ion Formation

In your textbook, read about chemical bonds and formation of ions.

Use each of the terms below just once to complete the passage.

chemical bond	electrons	energy level	ions	noble gases
nucleus	octet	pseudo-noble gas formations	valence	

The force that holds two atoms together is called a(n) (1) Chemical Bond.

Such an attachment may form by the attraction of the positively charged

(2) Nucleus of one atom for the negatively charged

(3) Electron of another atom, or by the attraction of charged atoms, which

are called (4) Ions. The attractions may also involve

(5) Valence electrons, which are the electrons in the outermost

(6) Energy level. The (7) Noble Gas are a family of

elements that have very little tendency to react. Most of these elements have a set of eight outermost

electrons, which is called a stable (8) Octet. The relatively stable electron

structures developed by loss of electrons in certain elements of groups 3, 4, 13, and 14 are called (9)

Pseudo-Noble gas formations.

For each statement below, write *true* or *false*.

- False 10. A positively charged ion is called an anion.
- True 11. Elements in group 1 lose their one valence electron, forming an ion with a 1+ charge.
- False 12. Elements tend to react so that they acquire the electron structure of a halogen.
- True 13. A sodium atom tends to lose one electron when it reacts.
- True 14. The electron structure of a zinc ion (Zn^{2+}) is an example of a pseudo-noble gas formation.
- False 15. A Cl^- ion is an example of a cation.
- True 16. The ending *-ide* is used to designate an anion.
- False 17. Nonmetals form a stable outer electron configuration by losing electrons and becoming anions.

Ionic Compounds And Metals Study Guide Answers

Kenneth W. Watkins



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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03

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Frentrup, 2002-06 Contains a brief overview of every chapter, review of skills, self-tests, and the answers and detailed solutions to all odd-numbered end-of-chapter problems in the text book. **Study Guide and Solutions Manual for**

Seager/Slabaugh's Chemistry for Today Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text. This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even-numbered end-of-chapter exercises. New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory

Application. **Study Guide and Solutions Manual for Fundamentals of General, Organic, and Biological Chemistry** Susan McMurry, 1992 *Study Guide and Partial Solutions Manual, Fundamentals of General, Organic, and Biological*

Chemistry Susan McMurry, 1996 This internationally acclaimed detective series is just the thing for lovers of those Number One Ladies looking for a darker, more realistic view of Botswana. Sue Baker Publishing News. **Student Study Guide and Solutions Manual to Accompany General, Organic, and Biochemistry** Katherine J. Denniston, Denniston Katherine, 2006

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon, buckyballs, graphite, and diamond, are illustrated at the left, as is the molecule methane, CH₄, from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane, found in natural gas, to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex, and the strength of this text is that it enables students with varied backgrounds to proceed together to

significant levels of achievement General, Organic, and Biological Chemistry Study Guide and Selected Solutions Karen C. Timberlake,2001-11 Keyed to the learning goals in the text this guide is designed to promote active learning through a variety of exercises with answers and mastery exams The guide also contains complete solutions to odd numbered problems

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