

7.05 Equilibrium Lab

Instructions: Complete all sections of this report and answer all questions in complete sentences for full credit. Use the [grading rubric](#) before you begin.

Click [here](#) to watch the [lab completion video](#). We show the entire lab simulation (so you can record the data, don't need to be logged into the course) and explain the conclusion/post lab.

Title: Investigating Equilibrium

Objective(s): Determine how increasing or decreasing the concentration of a reaction affects its equilibrium.

Hypothesis:

Complete the following statements for your hypotheses. Choose one of the options in brackets by **highlighting**, **bolding**, or placing your answer in a **different color**. (Careful: the first two are about the **REACTANTS** changing; the last two are about the **PRODUCTS** changing)

1. If the amount of reactant increases, the rate of the [**forward** or reverse] reaction will [**increase** or decrease] to reach a new equilibrium.
2. If the amount of reactant decreases, the rate of the [forward or **reverse**] reaction will [increase or **decrease**] to reach a new equilibrium.
3. If the amount of product increases, the rate of the [forward or **reverse**] reaction will [**increase** or decrease] to reach a new equilibrium.
4. If the amount of product decreases, the rate of the [**forward** or reverse] reaction will [increase or **decrease**] to reach a new equilibrium.

Introduction To Equilibrium Lab Answers

**Ralph Petrucci, William
Harwood, Geoffrey Herring**



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