

Ch. 5 Review

1. Matching: Match the following vocabulary terms to their correct definitions/descriptions.

A	state function		The ability to do work or transfer heat
B	heat		The energy a substance contains as a result of its chemical composition (the bonds within the compound)
C	exothermic		Quantitative measurement of the amount of heat within a substance
D	endothermic		A property of a system that is independent of the path taken and depends only on the final and initial conditions
E	calorimetry		Energy used to move an object (that has mass) over a distance
F	kinetic energy		A process which absorbs heat from its surroundings
G	potential energy		The amount of heat required to raise 1 gram of a substance 1°C
H	energy		The experimental measurement of the quantity of heat transferred during a physical or chemical reaction
I	work		A process which releases heat to its surroundings
J	temperature		Energy of motion that is directly tied to the temperature, or thermal energy, of a substance
K	specific heat capacity		Energy transferred as a result of differences in temperature

- The amount of heat (q) released or absorbed by a chemical reaction is equal to the enthalpy of the reaction (ΔH) if the _____ is held constant.
- What are the two conditions (of temperature & pressure) are required for a reaction to be considered a "standard state reaction"?
- What does the first Law of Thermodynamics state? How does this apply to the transfer of energy between a system and surroundings?

Modern Chemistry Ch 5 Review Answers

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