

Molarity

Molarity is just a unit of concentration that indicates how much solute we have in a solution. It is defined as shown below.

$$M = \frac{\text{mols of solute}}{\text{L. of solution}}$$

Thus, we need to remember the definitions of solute, solvent and solution:

solute: the substance dissolved in a solvent to form a solution; it is normally the component of a solution present in the smaller amount

solvent: the dissolving medium of a solution; it is normally the component of a solution present in the greater amount

solution: a mixture of substances that has a uniform composition; a homogeneous mixture

So, let's get some practice with this new unit!

Convert the following Molarity expressions into a conversion factor. The first one is done as an example.

1. $6.45 \text{ M NaCl} = \frac{6.45 \text{ mols NaCl}}{1 \text{ L. of NaCl sol'n}}$
2. $10.2 \text{ M HCl} = \underline{\hspace{2cm}}$
3. $3.42 \text{ M NaOH} = \underline{\hspace{2cm}}$
4. $0.0234 \text{ M NaHCO}_3 = \underline{\hspace{2cm}}$
5. $1.58 \times 10^{-9} \text{ M Na}_2\text{SO}_4 = \underline{\hspace{2cm}}$
6. $1.58 \times 10^{-4} \text{ M KCH}_3\text{CO}_2 = \underline{\hspace{2cm}}$

Now, let's try to do some calculations with Molarity! Remember to always write out M as the mol/L so you can see all the units.

7. How many moles of NaCl are in 2.50 liters of 6.45 M NaCl?

$$\frac{6.45 \text{ mols NaCl}}{1 \text{ L. NaCl}} \times 2.50 \text{ L. NaCl} = 16.1 \text{ mols NaCl}$$

8. How many moles of HCl are in 0.54 liters of 10.2 M HCl?
9. How many liters of NaOH are in 1.3 moles of 3.42 M NaOH?

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Hussin A.Rothana



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Table of Contents Molarity Calculations Answer Key Cavalcade With Work

1. Understanding the eBook Molarity Calculations Answer Key Cavalcade With Work
 - The Rise of Digital Reading Molarity Calculations Answer Key Cavalcade With Work
 - Advantages of eBooks Over Traditional Books
2. Identifying Molarity Calculations Answer Key Cavalcade With Work
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molarity Calculations Answer Key Cavalcade With Work
 - User-Friendly Interface
4. Exploring eBook Recommendations from Molarity Calculations Answer Key Cavalcade With Work

- Personalized Recommendations
 - Molarity Calculations Answer Key Cavalcade With Work User Reviews and Ratings
 - Molarity Calculations Answer Key Cavalcade With Work and Bestseller Lists
5. Accessing Molarity Calculations Answer Key Cavalcade With Work Free and Paid eBooks
- Molarity Calculations Answer Key Cavalcade With Work Public Domain eBooks
 - Molarity Calculations Answer Key Cavalcade With Work eBook Subscription Services
 - Molarity Calculations Answer Key Cavalcade With Work Budget-Friendly Options
6. Navigating Molarity Calculations Answer Key Cavalcade With Work eBook Formats
- ePub, PDF, MOBI, and More
 - Molarity Calculations Answer Key Cavalcade With Work Compatibility with Devices
 - Molarity Calculations Answer Key Cavalcade With Work Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Molarity Calculations Answer Key Cavalcade With Work
 - Highlighting and Note-Taking Molarity Calculations Answer Key Cavalcade With Work
 - Interactive Elements Molarity Calculations Answer Key Cavalcade With Work
8. Staying Engaged with Molarity Calculations Answer Key Cavalcade With Work
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Molarity Calculations Answer Key Cavalcade With Work
9. Balancing eBooks and Physical Books Molarity Calculations Answer Key Cavalcade With Work
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Molarity Calculations Answer Key Cavalcade With Work
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Molarity Calculations Answer Key Cavalcade With Work
- Setting Reading Goals Molarity Calculations Answer Key Cavalcade With Work
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Molarity Calculations Answer Key Cavalcade With Work

- Fact-Checking eBook Content of Molarity Calculations Answer Key Cavalcade With Work
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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