

## Molecular Polarity Lab

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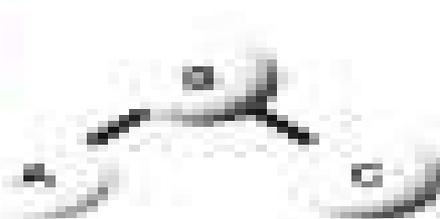
Circle the correct label above. Circle the right answer to the question.

### Part 1: Two atom molecules

1. Circle the two atoms that form dichloromethane and draw arrows to show electronegativity and the partial charges.  
ClC(Cl)Cl
2. Draw the electronegativity of **Cl** and **C**.  
Observe the arrow, partial charges and bond structure.  
Write the following information: the electronegativity of each of bonded atoms.  
Write:  
partial charges  
bond structure
3. Draw partial dipole arrows for a water molecule using the electronegativity of O and H and observe the results. Fill in the following information:
  - a. The polarity arrow points to the \_\_\_\_\_ electronegativity.
  - b. The partial positive charge is always on the \_\_\_\_\_ electronegative atom.
  - c. The larger the electronegativity difference the more \_\_\_\_\_ the bond structure.

### Part 2: Three atom molecules

1. Draw the central atom, electronegativity values and the electronegativity of **O** and **C**.  
CO2 Circle to show the bond dipole, molecular dipole and partial charges.
  - a. Draw the bond dipoles on the Lewis dot diagram.
  - b. In water and carbon dioxide partial molecular dipoles on the diagram.
  - c. Draw the partial charge symbols on the Lewis dot diagram.
  - d. Circle the electronegativity. Write the other electronegativity the molecule is composed of.
2. Circle the two atoms, CO2 showing the electronegativity values and the electronegativity of **O** and **C**.  
CO2 Circle to show the bond dipole, molecular dipole and partial charges.
  - a. Draw the bond dipoles on the Lewis dot diagram.
  - b. In water and carbon dioxide partial molecular dipoles on the diagram.
  - c. Draw the partial charge symbols on the Lewis dot diagram.
  - d. Circle the electronegativity. Write the other electronegativity the molecule is composed of.



# Molecule Polarity Phet Lab Answer Key

**ML Morrison**



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