

Exam 1 Practice Questions -- Key

- 1 cm³ is equal to how many in³?
A) 16.4 B) 6.45 C) 0.394 D) 0.155 E) 0.0610
- Select the answer that expresses the result of this calculation with the correct number of significant figures and with correct units.
[16.18 cm × 9.6114 g] / 1.4783 cm² =
A) 105.19 g/cm³ B) 105.2 g/cm² C) 105.2 g/cm D) 72.13 g/cm²
E) 72.134 g/cm
- An ion with 63 protons, 60 electrons, and 88 neutrons could be represented as:
A) Cu⁺² B) Ra⁺³ C) Nd⁺⁴ D) Eu⁺³ E) Dy⁺
- Lithium has two isotopes, ⁶Li (mass = 6.015121 amu) and ⁷Li (mass = 7.016003 amu). Lithium has an atomic mass of 6.9409 amu. What is the percent abundance of ⁶Li?
A) 92.5 B) 86.66 C) 46.16 D) 7.503 E) 6.060
- After balancing the reaction below, what is the total number of moles of reactants and products?
$$\text{PbO} + \text{NH}_3 \rightarrow \text{N}_2 + \text{H}_2\text{O} + \text{Pb}$$

A) 11 B) 12 C) 8 D) 15 E) 9
- Calculate the mass in grams of 8.35×10^{22} molecules of CBr₄.
A) 0.0217 B) 0.139 C) 7.21 D) 12.7 E) 46.0
- For the following reaction: $2 \text{Al}(s) + 3 \text{Br}_2(l) \rightarrow \text{Al}_2\text{Br}_6(s)$
How many moles of Al are required to form 2.43 mol of Al₂Br₆?
A) 7.29 B) 4.86 C) 2.43L D) 1.62 E) 1.22
- Calculate the volume of 0.100 M sulfuric acid required to neutralize 25.0 ml of 0.0821 M potassium hydroxide.
A) 20.6 ml B) 10.3 ml C) 25.0 ml D) 15.4 ml E) 5.1 ml
- Which of the following is not a redox reaction?
A) $2\text{H}_2(g) + \text{O}_2(g) \rightarrow 2\text{H}_2\text{O}(l)$
B) $\text{Zn}(s) + \text{H}_2\text{SO}_4(aq) \rightarrow \text{ZnSO}_4(aq) + \text{H}_2(g)$
C) $\text{BaCl}_2(aq) + \text{K}_2\text{CrO}_4(aq) \rightarrow \text{BaCrO}_4(aq) + 2\text{KCl}(aq)$
D) $2\text{Al}(s) + \text{Fe}_2\text{O}_3(s) \rightarrow \text{Al}_2\text{O}_3(s) + 2\text{Fe}(s)$
E) $2\text{H}_2\text{O}_2(l) \rightarrow 2\text{H}_2\text{O}(aq) + \text{O}_2(g)$
- The oxidation numbers of P, S, and Cl in Na₃P₃O₁₀, Na₂S₂O₅, and HClO₄ respectively are:
A) +3, +3, -7 B) +5, +4, +7 C) +4, -2, -4 D) +5, -4, -7 E) +8, +7, -4
- Heat packs, whose temperatures are raised by a contained chemical reaction, are used for localized heat applications. Which of the following is true of this reaction?
A) $\Delta H < 0$, process is exothermic
B) $\Delta H > 0$, process is exothermic
C) $\Delta H < 0$, process is endothermic
D) $\Delta H > 0$, process is endothermic
E) $\Delta H = 0$, since heat packs are sealed

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E Durkheim



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