

Mayya's Equilibrium Study Guide



K_c is the Equilibrium Constant

$$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$



K_p is the Equilibrium Constant
when pressures are given

$$K_p = \frac{P_c^c P_d^d}{P_a^a P_b^b}$$

$$K_p = K_c(RT)^{\Delta n}$$

Where $R=0.08206 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K}$, T = temperature in K,
 Δn = moles of gas products - moles of gas reactants

How To Predict The Direction Of Reaction

$$Q = \frac{[C]^c \text{ noneq} [D]^d \text{ noneq}}{[A]^a \text{ noneq} [B]^b \text{ noneq}}$$

If $Q > K$, the reaction will go to the left
If $Q < K$, the reaction will go to the right
If $Q = K$, the reaction is at equilibrium

ICE CHART

I = Initial (initial concentrations given in the problem)
C = Change (in terms of X, including coefficients)
E = Equilibrium (I+C)

Make an ICE Chart when you are asked to find equilibrium concentrations or pressures

Equation Stoichiometry & Equilibrium Constant

- Reverse an equation = Take reciprocal of K ($1/K$)
- Multiply an equation by a constant = Raise K to the power of that constant
- Add equations = Multiply their K 's

Equilibrium And Ksp Study Guide

Pamela Mills, Amina El-Ashmawy



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web the work to be covered in each of the grades 10 11 and 12 levels has been described separately the aims and objectives of teaching mathematics at senior secondary school level have been derived from three sources the educational reform document of 1997 the structure of mathematics as an academic discipline and the needs of the society