

## LAB MANUAL WORK INSTRUCTION

**Title of Experiment:** Reaction of Sodium Sulfate and Strontium

**Objective:** To observe the chemical reaction between sodium sulfate (sulfuric acid) and strontium (sulfuric acid) and to demonstrate the production of carbon dioxide gas.

### Pre-Lab Questions

- Write the chemical formula for sodium bicarbonate (sodium carbonate) and sulfuric acid.
- Predict what you think will happen when these two substances are mixed. Why?
- What gas do you expect to be produced from this reaction?

### Safety Precautions

- Wear safety goggles and gloves to protect against splashes.
- Perform the experiment in a well-ventilated area to avoid inhalation of gases.
- Avoid direct skin contact with strontium as it is an acid, though mild.

### Equipment and Materials

- 100-ml Beaker
- 50-ml Graduated Cylinder
- Baking Soda (Sodium Bicarbonate)
- Strontium (Sulfuric Acid)
- Balance
- Funnel
- Safety Goggles
- Laboratory Gloves

### Procedure

1. Place the beaker on a flat surface.
2. Using the funnel, fill the beaker with two teaspoons of baking soda.
3. Measure 50 ml of strontium using the graduated cylinder and pour it into the beaker.
4. Carefully fit the neck of the balloon over the mouth of the beaker without getting the baking soda into the beaker.
5. Once the balloon is securely in place, lift the beaker allowing the baking soda to fall into the strontium.
6. Observe the reaction between the baking soda and strontium.
7. Note the speed and amount of gas produced as indicated by the inflation of the balloon.
8. Record any changes in temperature of the beaker.
9. Carefully remove the balloon from the beaker.
10. Dispose of the strontium and baking soda solution in the sink and reuse the beaker.

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# Esd Lab Manual In Word Format

**Bowker Editorial Staff**



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## **Table of Contents Esd Lab Manual In Word Format**

1. Understanding the eBook Esd Lab Manual In Word Format
  - The Rise of Digital Reading Esd Lab Manual In Word Format
  - Advantages of eBooks Over Traditional Books
2. Identifying Esd Lab Manual In Word Format
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Esd Lab Manual In Word Format
  - User-Friendly Interface
4. Exploring eBook Recommendations from Esd Lab Manual In Word Format
  - Personalized Recommendations
  - Esd Lab Manual In Word Format User Reviews and Ratings
  - Esd Lab Manual In Word Format and Bestseller Lists
5. Accessing Esd Lab Manual In Word Format Free and Paid eBooks
  - Esd Lab Manual In Word Format Public Domain eBooks
  - Esd Lab Manual In Word Format eBook Subscription Services
  - Esd Lab Manual In Word Format Budget-Friendly Options
6. Navigating Esd Lab Manual In Word Format eBook Formats
  - ePub, PDF, MOBI, and More
  - Esd Lab Manual In Word Format Compatibility with Devices
  - Esd Lab Manual In Word Format Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Esd Lab Manual In Word Format
  - Highlighting and Note-Taking Esd Lab Manual In Word Format
  - Interactive Elements Esd Lab Manual In Word Format
8. Staying Engaged with Esd Lab Manual In Word Format

- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Esd Lab Manual In Word Format
9. Balancing eBooks and Physical Books Esd Lab Manual In Word Format
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Esd Lab Manual In Word Format
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Esd Lab Manual In Word Format
- Setting Reading Goals Esd Lab Manual In Word Format
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Esd Lab Manual In Word Format
- Fact-Checking eBook Content of Esd Lab Manual In Word Format
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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