

Key

# Molecular Geometry

How can molecular shapes be predicted using the VSEPR theory?

## Why?

When you draw a Lewis structure for a molecule on paper, you are making a two-dimensional representation of the atoms. In reality however, molecules are not flat—they are three-dimensional. The true shape of a molecule is important because it determines many physical and chemical properties for the substance. In this activity you will learn how to predict molecular shapes.

## Model 1 – Lewis Structures

Lewis Structures	Electron Domains	3-D Molecular Shape
1. $H_2CO$ 	$H_2CO$ 3 electron domains (3 bonding, 0 nonbonding)	
2. $PF_3$ 	$PF_3$ 4 electron domains (3 bonding, 0 nonbonding)	
3. $CH_4$ 	$CH_4$ 4 electron domains (4 bonding, 0 nonbonding)	
4. $NH_3$ 	$NH_3$ 4 electron domains (3 bonding, 1 nonbonding)	
5. $H_2O$ 	$H_2O$ 4 electron domains (2 bonding, 2 nonbonding)	
6. $CO_2$ 	$CO_2$ 2 electron domains (2 bonding, 0 nonbonding)	

Lone pair = ..

# Geometry Pogil Answer Key

**Roman Wölfel**



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