



Name:

Date:

Student Exploration: Calorimetry Lab

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: calorie, calorimeter, joule, specific heat capacity

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. The Latin word *calor* means "heat," and *meter* comes from the Greek word meaning "to measure." What do you think a **calorimeter** does?

measuring the heat of chemical reactions or physical changes as well as heat capacity

2. Where have you heard the word **calorie** before? What do you think a calorie is?

calorie, a unit of energy or heat variously defined. the calorie was originally defined as the amount of heat required at a pressure of 1 standard atmosphere to raise the temperature of 1 gram of water 1 degree celsius.

Gizmo Warm-up

A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up.

Calorimeters can be used to find a substance's **specific heat capacity**. You will use the *Calorimetry Lab* Gizmo to determine the specific heat capacities of various substances.



1. On the SIMULATION pane, select **Copper**. Use the slider to set its **Mass** to 200 g. Set the **Water mass** to 200 g. Check that the **Water temp** is set to 30.0 °C and the copper's **Temp** is 90 °C. Select the GRAPH tab, and click **Play** (▶).

- A. What was the **Final temperature** of the copper and the water?
- B. How much did the temperature of the copper change?
- C. How much did the temperature of the water change?

34.96

down 55.04

down 165.04

2. Specific heat capacity can be described as a substance's resistance to temperature changes. Which substance has a greater specific heat capacity, copper or water? Explain.

Fan Cart Gizmo Answer Key

Guillaume Favre



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Table of Contents Fan Cart Gizmo Answer Key

1. Understanding the eBook Fan Cart Gizmo Answer Key
 - The Rise of Digital Reading Fan Cart Gizmo Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Fan Cart Gizmo Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Fan Cart Gizmo Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Fan Cart Gizmo Answer Key
 - Personalized Recommendations
 - Fan Cart Gizmo Answer Key User Reviews and Ratings
 - Fan Cart Gizmo Answer Key and Bestseller Lists
5. Accessing Fan Cart Gizmo Answer Key Free and Paid eBooks
 - Fan Cart Gizmo Answer Key Public Domain eBooks
 - Fan Cart Gizmo Answer Key eBook Subscription Services
 - Fan Cart Gizmo Answer Key Budget-Friendly Options

6. Navigating Fan Cart Gizmo Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Fan Cart Gizmo Answer Key Compatibility with Devices
 - Fan Cart Gizmo Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Fan Cart Gizmo Answer Key
 - Highlighting and Note-Taking Fan Cart Gizmo Answer Key
 - Interactive Elements Fan Cart Gizmo Answer Key
8. Staying Engaged with Fan Cart Gizmo Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Fan Cart Gizmo Answer Key
9. Balancing eBooks and Physical Books Fan Cart Gizmo Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Fan Cart Gizmo Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Fan Cart Gizmo Answer Key
 - Setting Reading Goals Fan Cart Gizmo Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Fan Cart Gizmo Answer Key
 - Fact-Checking eBook Content of Fan Cart Gizmo Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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