

SENIOR FIVE

- 1(a). What is meant by the term pH  
(b). A solution of magnesium hydroxide contains 0.0002 moles per litre of magnesium ions.  
(i). Calculate the number of moles of hydroxide ions present in the solution (ii).  
The pH of the solution. ( $K_w = 1.0 \times 10^{-14}$ )

2. Define the term boiling point elevation constant.  
(b) Describe an experiment to show how molecular mass of urea can be determined using ebulyoscopy.  
(c) Acetone boils at  $56.38^\circ\text{C}$  and a solution of 1.41g of an organic solid in 20g of a cetone boils at  $56.88^\circ\text{C}$ . Calculate the mass of one mole of the organic solid (kb for a cetone per 100g is 16.7)  
(d) (i) Explain why freezing point depression and boiling point elevation method are not suitable for determination of molar mass of polymers

- 3.(a) Define the term pH  
(b) Find out the pH of 0.002M acetic acid solution if it is 2.3% Ionised in dilute solution  
(c) (i) Explain the meaning of the word "Buffer solution"  
(ii) Explain the mechanism of action of a basic buffer.  
(d) (i) The  $\text{P}^{\text{H}}$  of a solution containing 0.15Mol per litre ethanoic acid and a certain mass of sodium ethanoate is found to be 4.8697. Calculate the mass of sodium ethanoate in the solution (ka for ethanoic acid is  $1.8 \times 10^{-5}$ )  
(ii) Calculate the PH change that would occur when  $1\text{cm}^3$  if 1M HCl is added to the above solution.  
(e) A chemist needs a buffered solution of propanoic acid and sodium propanoate.  
Calculate the ratio of  $(\text{CH}_3\text{CH}_2\text{COOH})/(\text{CH}_3\text{CH}_2\text{COONa})$  required to yield a  $\text{p}^{\text{H}}$  of 4.3 (Ka for propanoic acid is  $1.3 \times 10^{-5}$ )

4. Explain four applications of butter solutions. (a) Define the following terms  
(i) Atomic radius (iii) Electron affinity  
(ii) Ionization energy (iv) Electronegativity  
(b) Explain the variation of the following periodic properties across and down the group.

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