Name:_____ Date:_____ Period:_____

# Molar Mass Practice L# – *Answer Key*

**Objective:** - Calculate molar mass using the periodic table

**Directions:** Find the molar mass of the following atoms, molecules, or formula units. Round answer to the hundredths place. Don't forget your units! You must show your work!

1. Ti     [47.88 g/mol]

2. Fe     [55.85 g/mol]

3. $CO_2$
C: 1 x 12.01 =     12.01
O: 2 x 16.00 =     32.00
    [44.01 g/mol]

4. LiCl
Li: 1 x 6.94 =      6.94
Cl: 1 x 35.45 =     35.45
    [42.39 g/mol]

5. $O_2$
O: 2 x 16.00 = 32.00
    [32.00 g/mol]

6. Iron(III) oxide     $Fe_2O_3$

(2×55.85) + (3×16.00) = [159.7 g/mol]

7. $H_2SiO_3$

    [78.10 g/mol]

8. Lead(II) hydroxide     $Pb(OH)_2$

    [241.2 g/mol]

9. $(NH_4)_2S$

(2×14.01) + (8×1.01) + 32.07 = [68.17 g/mol]

10. $C_2H_5OH$
C: 2 x 12.01 = 24.02
H: 6 x 1.01 =   6.06
O: 1 x 16.00 = 16.00
    [46.08 g/mol]

11. $Mg(NO_2)_2$
Mg: 1 x 24.31 = 24.31
N: 2 x 14.01 = 28.02
O: 4 x 16.00 = 64.00
    [116.33 g/mol]

12. $H_2SO_4$

(2×1.01) + 32.07 + (4×16.00) = [98.09 g/mol]

13. $Mn_3(PO_4)_2$
Mn: 3 x 54.94 = 164.82
P: 2 x 30.97 = 61.94
O: 8 x 16.00 = 128.00
    [354.76 g/mol]

14. Dihydrogen monoxide     $H_2O$

(2×1.01) + 16.00 = [18.02 g/mol]

15. $Zn(NO_3)_2$

    [189.4 g/mol]

16. Aluminum sulfate     $Al_2(SO_4)_3$

    [342.2 g/mol]

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