

## Ch 5 Gas Law Notes

### Pressure Problems:

1. How many millimeters of mercury is equal to 1.50 atmosphere?

**1140 mmHg**

2. Convert the pressure of 955 mm of Hg to the following units:

a. atmospheres **1.26 atm**

b. kilopascals **127 kPa**

c. pascals **127000 Pa**

3. If a force of 9.0 Newtons is applied to an area of  $2.0 \text{ m}^2$ , what is the pressure?

**4.5 Pa**

4. An open ended manometer measures the atmospheric pressure to be 102 kPa. The difference in the height of the mercury is 30 mm Hg. What is the pressure exerted by the gas?

**795 mmHg**

5. A sample of a gas under a pressure of 822 kPa has a volume of  $312 \text{ cm}^3$ . The pressure is increased to a 948 kPa. What is the new volume of the gas assuming that the temperature is constant? State the gas law that is applied to this problem \_\_\_\_\_.

**270.  $\text{cm}^3$ ; Boyle's Law**

6. The volume of a gas is  $204 \text{ cm}^3$  when the pressure is 925 kPa. A change in pressure caused the volume to change to  $306 \text{ cm}^3$ . What is the pressure that causes this change in volume? State the gas law \_\_\_\_\_.

**616 kPa; Boyle's Law**

7. A sample of gas has a volume of  $152 \text{ cm}^3$  when its temperature is  $18^\circ\text{C}$ . If the temperature is increased to  $32^\circ\text{C}$ , what will the new volume be assuming that the pressure is constant? State the gas law \_\_\_\_\_.

**159  $\text{cm}^3$ ; Charles's Law**

8. A sample of gas has a volume of  $125 \text{ cm}^3$  when its temperature is  $25^\circ\text{C}$ . If the volume changes to  $98 \text{ cm}^3$ , what is the new temperature? State the gas law \_\_\_\_\_.

**230K; Charles's Law**

9. A gas has a volume of  $31.2 \text{ dm}^3$  at a temperature of  $28^\circ\text{C}$  and a pressure of 82.6 kPa. How many moles are there in the gas sample? How many molecules are in the sample? The gas is an ideal gas. State the law \_\_\_\_\_.

**1.03 mol;  $6.19 \times 10^{23}$  molecules; ideal gas law**

10. What volume will 15.0 g of carbon monoxide (CO) occupy at STP? The gas is an ideal gas. State the law \_\_\_\_\_.

**12.0L; ideal gas law**

# Gas Law Review Sheet Answers

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