

Practice with the Gas Laws with Answers

All temperatures must be in **KELVIN** STP = **0 °C and 1.0 atm**

Converting one unit of pressure to another

1. Change 758. mm Hg to Torr

$$758 \text{ mm Hg} \times \frac{760 \text{ Torr}}{760 \text{ mm Hg}} = 758 \text{ Torr}$$

2. Change 771.3 mm Hg to atm

$$771.3 \text{ mm Hg} \times \frac{1 \text{ atm}}{760 \text{ mm Hg}} = 1.015 \text{ atm}$$

3. Change 95.6 kPa to mm Hg

$$95.6 \text{ kPa} \times \frac{760 \text{ mm Hg}}{101.325 \text{ kPa}} = 717 \text{ mm Hg}$$

4. Is the pressure of the gas sample greater or less than the pressure of the atmosphere? Explain.

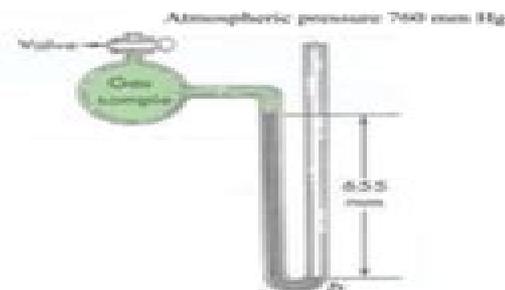
Less, the mercury line is higher on that side

What is the pressure of the gas in this manometer (in mm Hg)?

$$\begin{array}{r} 760 \text{ mm Hg} \\ -655 \text{ mm Hg} \\ \hline 105 \text{ mm Hg} \end{array}$$

Convert that pressure to kPa.

$$105 \text{ mm Hg} \times \frac{101.325 \text{ kPa}}{760 \text{ mm Hg}} = 14.0 \text{ kPa}$$



Calculating Density of a Gas at STP

1. What is the density of dinitrogen tetroxide at STP?

$$\text{N}_2\text{O}_4 \quad \frac{92.01 \text{ g}}{22.4 \text{ L}} = 4.108 \text{ g/L}$$

2. What is the density of sulfur trioxide at STP?

$$\text{SO}_3 \quad \frac{80.06 \text{ g}}{22.4 \text{ L}} = 3.574 \text{ g/L}$$

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review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields

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