

Gas Variables

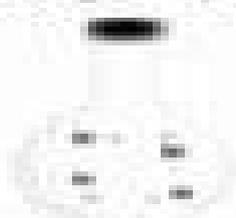
How are the variables that describe a gas related?

Why?

Imagine buying a balloon bouquet at a party store. How will the helium gas in the bouquet behave if you carry it outside on a hot summer day? How will it behave if you carry it outside during a snowstorm? What happens if the balloons are made of latex, which can stretch? What happens if the balloons are made of Mylar[®], which cannot stretch? What if you add just a small amount of gas to each balloon? What if you add a lot of gas? In this activity, you will explore four variables that quantify gases—pressure (P), volume (V), temperature (T), and amount (n) of gas. These four variables can be related mathematically so that predictions about gas behavior can be made.

Model 1 – Gases in a Nonflexible Container

Experiment A (Adding more gas)



A1

Volume = 2 units

External pressure = 1 atm Internal pressure = 1 atm External pressure = 1 atm
Internal pressure = 1 atm Internal pressure = 1 atm Internal pressure = 1 atm
Temperature = 300 K Temperature = 300 K Temperature = 300 K



A2

Volume = 2 units

External pressure = 1 atm Internal pressure = 2 atm External pressure = 1 atm
Internal pressure = 1 atm Internal pressure = 1 atm Internal pressure = 1 atm
Temperature = 300 K Temperature = 300 K Temperature = 300 K



A3

Volume = 2 units

External pressure = 1 atm Internal pressure = 3 atm External pressure = 1 atm
Internal pressure = 1 atm Internal pressure = 1 atm Internal pressure = 1 atm
Temperature = 300 K Temperature = 300 K Temperature = 300 K

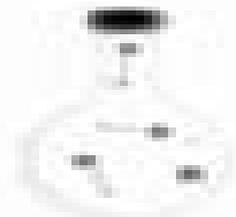
Experiment B (Heating the gas)



B1

Volume = 2 units

External pressure = 1 atm Internal pressure = 1 atm External pressure = 1 atm
Internal pressure = 1 atm Internal pressure = 1 atm Internal pressure = 1 atm
Temperature = 300 K Temperature = 300 K Temperature = 300 K



B2

Volume = 2 units

External pressure = 1 atm Internal pressure = 2 atm External pressure = 1 atm
Internal pressure = 1 atm Internal pressure = 1 atm Internal pressure = 1 atm
Temperature = 400 K Temperature = 400 K Temperature = 400 K



B3

Volume = 2 units

External pressure = 1 atm Internal pressure = 3 atm External pressure = 1 atm
Internal pressure = 1 atm Internal pressure = 1 atm Internal pressure = 1 atm
Temperature = 600 K Temperature = 600 K Temperature = 600 K

Note: Volume is the model's way of representing the fixed volume of gas at the conditions given. Each cube represents 1×10^{23} molecules. These numbers are much larger compared to the space between them than in a liquid or a solid.

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TD Snyder



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Study Guide for Fundamentals of Nursing Carol Lillis, Priscilla LeMone, Marilee LeBon, Pamela Lynn, 2010-11-01 Created in conjunction with *Fundamentals of Nursing Seventh Edition* this Study Guide helps students review and apply concepts from the textbook to prepare for exams as well as nursing practice Each chapter includes three sections Practicing for NCLEX containing multiple choice and alternate format questions Developing Your Knowledge Base including a variety of questions formats such as fill in the blank matching and short answer and Applying Your Knowledge comprised of critical thinking questions reflective practice scenarios and patient care studies An Answer Key appears at the back of the book

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In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Gas Variables Packet Answers Pg 38**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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