

Chapter 5: Gases - Things to know*

Please remember that simply knowing or memorizing the information below is not sufficient. You must be able to know when and how to APPLY the items below appropriately.

5.1 Basics

Pressure and Units *(see guide!)*

5.2 Simple Gas Laws *(see guide!)*

5.3 Ideal Gas Law *(see preview below!)*

Used to determine one gas variable for a given scenario, knowing the other variables.

A. $PV=nRT$ ($R = 0.0821 \text{ L atm/mol K}$; $8.314 \text{ L kPa/mol K}$; be sure to match all units to the constant)

B. Limitations:

1. Works well at low pressures *as the molecules of the gas rarely collide or come into close enough proximity for intermolecular forces to be significant*
2. Works well at high temperatures *as the molecules of the gas have so much energy that the intermolecular forces and energy lost in collisions is negligible*
3. Most gases do not behave ideally above 1 atm pressure
4. Does not work well near the condensation conditions of a gas
5. Deviations occur due to stronger attractive forces

C. Derivation to relate molar mass and density

$D = m/V$ $n = m/M$ where M is molar mass

$PV = m/MRT$; $M = DRT/P$

5.4 Gas Mixtures and Collecting Gas of Water

5.5 Diffusion and Effusion

Gases Study Guide

M Lipman



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Table of Contents Gases Study Guide

1. Understanding the eBook Gases Study Guide
 - The Rise of Digital Reading Gases Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Gases Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gases Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gases Study Guide
 - Personalized Recommendations
 - Gases Study Guide User Reviews and Ratings
 - Gases Study Guide and Bestseller Lists

5. Accessing Gases Study Guide Free and Paid eBooks
 - Gases Study Guide Public Domain eBooks
 - Gases Study Guide eBook Subscription Services
 - Gases Study Guide Budget-Friendly Options
6. Navigating Gases Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Gases Study Guide Compatibility with Devices
 - Gases Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gases Study Guide
 - Highlighting and Note-Taking Gases Study Guide
 - Interactive Elements Gases Study Guide
8. Staying Engaged with Gases Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gases Study Guide
9. Balancing eBooks and Physical Books Gases Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gases Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gases Study Guide
 - Setting Reading Goals Gases Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gases Study Guide
 - Fact-Checking eBook Content of Gases Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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