

Complete the following calculation by list the given information, rewriting the formula to solve for the unknown, and plugging in the information (number and units), and writing the answer significantly.

1. A sample of 0.500 moles of gas is placed in a container of volume of 2.50 L. What is the pressure of the gas in torr if the gas is at 25 °C? (Ideal Gas Law)

$$P = \frac{nRT}{V} = \frac{(0.500 \text{ moles}) (0.0821 \text{ L} \cdot \text{atm}) (298 \text{ K})}{(\text{mole} \cdot \text{K}) (2.50 \text{ L})} = P = 4.89 \text{ atm} = 3720 \text{ torr}$$

2. If 2.00 mol of gas occupies 4.50 L at STP. How much of the same gas will occupy 3.00 L at STP? (Avogadro's Law)

$$n_2 = \frac{V_2 n_1}{V_1} = \frac{(3.00 \text{ L}) (2.00 \text{ moles})}{(4.50 \text{ L})} = 1.33 \text{ moles}$$

3. Determine the partial pressures of each of the gases in the following mixture: 17.04 g NH<sub>3</sub>, 40.36 g Ne and 19.00 g F<sub>2</sub>. The gases are at 1.50 atm of pressure. (Dalton's Law of Partial Pressure & Mole Ratios)

$$\frac{17.04 \text{ g NH}_3}{17.04 \text{ g NH}_3} \left| \frac{1 \text{ mole NH}_3}{17.04 \text{ g NH}_3} \right| \left| \frac{1.000 \text{ mole NH}_3}{3.500 \text{ moles}} \right| \quad \frac{1.000 \text{ mole NH}_3}{3.500 \text{ moles}} = 0.2857 \times 1.50 \text{ atm} = 0.429 \text{ atm NH}_3$$

$$\frac{40.36 \text{ g Ne}}{20.18 \text{ g Ne}} \left| \frac{1 \text{ mole Ne}}{20.18 \text{ g Ne}} \right| \left| \frac{2.000 \text{ moles Ne}}{3.500 \text{ moles}} \right| \quad \frac{2.000 \text{ moles Ne}}{3.500 \text{ moles}} = 0.5714 \times 1.50 \text{ atm} = 0.857 \text{ atm Ne}$$

$$\frac{19.00 \text{ g F}_2}{38.00 \text{ g F}_2} \left| \frac{1 \text{ mole F}_2}{38.00 \text{ g F}_2} \right| \left| \frac{0.5000 \text{ moles F}_2}{3.500 \text{ moles}} \right| \quad \frac{0.5000 \text{ moles F}_2}{3.500 \text{ moles}} = 0.1429 \times 1.50 \text{ atm} = 0.214 \text{ atm F}_2$$

$$1.000 \text{ moles} + 2.000 \text{ moles} + 0.5000 \text{ moles} = 3.500 \text{ moles}$$

4. If at 1.00 atm of pressure water boils at 100. °C, at what temperature would water boil if the pressure is 600. torr? (This shows why food doesn't cook well at higher elevations) (Gay-Lussac's Law)

$$T_2 = \frac{T_1 P_2}{P_1} = \frac{(373 \text{ K}) (0.789 \text{ atm})}{(1.00 \text{ atm})} = 294 \text{ K}$$

5. Calculate the volume of 40.6 g of F<sub>2</sub> at STP. (Ideal Gas Law)

$$\frac{40.6 \text{ g F}_2}{38.00 \text{ g F}_2} \left| \frac{1 \text{ mole F}_2}{38.00 \text{ g F}_2} \right| \left| \frac{22.4 \text{ L F}_2}{1 \text{ mole F}_2} \right| \left| \frac{23.9 \text{ L F}_2}{22.4 \text{ L F}_2} \right|$$

or

$$\frac{40.6 \text{ g F}_2}{38.00 \text{ g F}_2} \left| \frac{1 \text{ mole F}_2}{38.00 \text{ g F}_2} \right| \left| \frac{1.07 \text{ mole F}_2}{1 \text{ mole F}_2} \right|$$

$$V = \frac{nRT}{P} = \frac{(1.07 \text{ moles}) (0.0821 \text{ L} \cdot \text{atm}) (273 \text{ K})}{(\text{mole} \cdot \text{K}) (1.00 \text{ atm})} = 24.0 \text{ L F}_2$$

6. If the winter pressure that most car tire tires should be at to wear evenly is 32.0 psi, what is the pressure in atm? (Pressure Conversion)

$$\frac{32.0 \text{ psi}}{14.7 \text{ psi}} \left| \frac{1.00 \text{ atm}}{14.7 \text{ psi}} \right| \left| \frac{2.18 \text{ atm}}{1.00 \text{ atm}} \right|$$

# Gases Chem Answer Key

**Michael Seilmaier**



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## **Table of Contents Gases Chem Answer Key**

1. Understanding the eBook Gases Chem Answer Key
  - The Rise of Digital Reading Gases Chem Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Gases Chem Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gases Chem Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gases Chem Answer Key
  - Personalized Recommendations
  - Gases Chem Answer Key User Reviews and Ratings
  - Gases Chem Answer Key and Bestseller Lists
5. Accessing Gases Chem Answer Key Free and Paid eBooks
  - Gases Chem Answer Key Public Domain eBooks
  - Gases Chem Answer Key eBook Subscription Services
  - Gases Chem Answer Key Budget-Friendly Options

6. Navigating Gases Chem Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Gases Chem Answer Key Compatibility with Devices
  - Gases Chem Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gases Chem Answer Key
  - Highlighting and Note-Taking Gases Chem Answer Key
  - Interactive Elements Gases Chem Answer Key
8. Staying Engaged with Gases Chem Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Gases Chem Answer Key
9. Balancing eBooks and Physical Books Gases Chem Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gases Chem Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Gases Chem Answer Key
  - Setting Reading Goals Gases Chem Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gases Chem Answer Key
  - Fact-Checking eBook Content of Gases Chem Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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