

## Stoichiometry Practice 1.1 - Answer Key

**Objectives:** - Use stoichiometry to calculate the mass of any substance involved in a chemical reaction.

**Directions:** Answer the following stoichiometry problems and show your work.

Mass A → Mole A → Mole B:

1. When calcium hydroxide and phosphoric acid ( $\text{H}_3\text{PO}_4$ ) are combined, calcium phosphate and water are produced. If the reaction starts with 10.0 grams of phosphoric acid, how many moles of water are produced?



$$10.0 \text{ g} = \left( \frac{2 \text{ mol H}_3\text{PO}_4}{97.99 \text{ g H}_3\text{PO}_4} \right) \left( \frac{6 \text{ mol H}_2\text{O}}{2 \text{ mol H}_3\text{PO}_4} \right) = \boxed{0.308 \text{ mol H}_2\text{O}}$$

Mole A → Mole B → Mass B:

2. How many grams of oxygen are needed in the combustion 1.07 mol of methyl alcohol ( $\text{CH}_3\text{OH}$ )?



$$1.07 \text{ mol} = \left( \frac{3 \text{ mol O}_2}{2 \text{ mol CH}_3\text{OH}} \right) \left( \frac{32.00 \text{ g O}_2}{1 \text{ mol O}_2} \right) = \boxed{51.4 \text{ g}}$$

Mass A → Mole A → Mole B → Mass B:

3. How many grams of copper are produced when 3.46 grams of aluminum react with excess copper (II) acetate solution?  $\boxed{13.2 \text{ g}}$



$$3.46 \text{ g} = \left( \frac{2 \text{ mol Al}}{26.98 \text{ g Al}} \right) \left( \frac{3 \text{ mol Cu}}{2 \text{ mol Al}} \right) \left( \frac{63.54 \text{ g Cu}}{1 \text{ mol Cu}} \right) = \boxed{13.2 \text{ g Cu}}$$

4. Lithium sulfate and barium chloride react to produce barium sulfate and lithium chloride. How many grams of barium chloride are needed to react with 17.5 g of lithium sulfate?



$$17.5 \text{ g} = \left( \frac{1 \text{ mol Li}_2\text{SO}_4}{148.98 \text{ g Li}_2\text{SO}_4} \right) \left( \frac{1 \text{ mol BaCl}_2}{1 \text{ mol Li}_2\text{SO}_4} \right) \left( \frac{208.23 \text{ g BaCl}_2}{1 \text{ mol BaCl}_2} \right) = \boxed{24.1 \text{ g}}$$

# Mass Mass Stoichiometry Practice Key

**DP Hallahan**



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