

Chapter 3 Mass Relationships in Chemical Reactions

Student: _____

- An atom of bromine has a mass about four times greater than that of an atom of neon. Which choice makes the correct comparison of the relative numbers of bromine and neon atoms in 1,000 g of each element?
 - The number of bromine and neon atoms is the same.
 - There are one thousand times as many bromine atoms as neon atoms.
 - There are one thousand times as many neon atoms as bromine atoms.
 - There are four times as many neon atoms as bromine atoms.
 - There are four times as many bromine atoms as neon atoms.
- An atom of bromine has a mass about four times greater than that of an atom of neon. How many grams of neon will contain the same number of atoms as 1,000 g of bromine?
 - 4 g Ne
 - 250 g Ne
 - 400 g Ne
 - 1,000 g Ne
 - 4,000 g Ne
- What is the average mass, in grams, of one atom of iron?
 - 6.02×10^{23} g
 - 1.66×10^{-24} g
 - 9.28×10^{-23} g
 - 55.85 g
 - 55.85×10^{-23} g
- What is the mass, in grams, of one arsenic atom?
 - 5.48×10^{-23} g
 - 33.0 g
 - 74.9 g
 - 1.24×10^{-22} g
 - 8.04×10^{21} g
- What is the mass, in grams, of one copper atom?
 - 1.055×10^{-22} g
 - 63.55 g
 - 1 amu
 - 1.66×10^{-24} g
 - 9.476×10^{-22} g

Mass Relationships In Chemical Reactions Practice Test

JL Elias



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