

## REDOX REACTIONS.

### NOTES...

#### Intro

~Many Chemical reactions involve transfer of electrons from one chemical substance to another. These electron-transfer reactions are termed as Oxidation-Reduction or Redox reactions.

#### Oxidation

(i) Loss of electron

- Example:  $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^{-}$

- Sub-points:

- Loss of hydrogen
- Gain of oxygen
- Gain of **electronegative** element OR loss of electropositive element

#### Reduction

i) Gain of electron

- Example:  $\text{Cl}_2 + 2\text{e}^{-} \rightarrow 2\text{Cl}^{-}$

- Sub-points:

- Gain of hydrogen
- Loss of oxygen
- Loss of electronegative element OR gain of electropositive element

#### Redox reaction

The reaction in which oxidation as well as reduction takes place together is known as a redox reaction.

#### ~REDOX REACTIONS IN TERMS OF ELECTRON TRANSFER (MODERN IDEA)

Oxidation process is the process in which one or more electrons are lost by an atom, ion, or molecule.

Example:  $\text{Zn} \rightarrow \text{Zn}^{2+} + 2\text{e}^{-}$

# Mcgraw Hill Oxidation Reduction Study Guide

**Padhraic Smyth**



**Mcgraw Hill Oxidation Reduction Study Guide:**

## Reviewing **Mcgraw Hill Oxidation Reduction Study Guide**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is really astonishing. Within the pages of "**Mcgraw Hill Oxidation Reduction Study Guide**," an enthralling opus penned by a very acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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