

Name: _____ Class: _____ Date: _____

Feel the Heat!

Part 1: Temperature Changes $Q = mC\Delta T$

- 1.) 100.0 mL of 4°C water is heated until it reaches 37°C. Assume that the density of liquid water is 1.0 g/mL and that the specific heat capacity of water is 4.18 J/g°C. What amount of heat energy was added to cause this raise in temperature?

$$Q = (100 \text{ g})(4.184)(37-4) = 13,807 \text{ J}$$

- 2.) 100.0 mL of water is cooled from 75°C to 25°C. Assume that the density of liquid water is 1.0 g/mL and that the specific heat capacity of water is 4.18 J/g°C. What amount of heat energy was lost to cause this raise in temperature?

$$Q = (100 \text{ g})(4.184)(25-75) = -20,920 \text{ J (heat lost)}$$

- 3.) How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C? The specific heat capacity of aluminum is 0.90 J/g°C.

$$Q = (10 \text{ g})(0.90)(55-22) = 297 \text{ J}$$

- 4.) Calculate the specific heat capacity for wood if 1500. g of the wood absorbs 6.75×10^4 Joules of heat and its temperature changes from 32°C to 57°C

$$6.75 \times 10^4 \text{ J} = (1500 \text{ g}) C (57-32)$$

$$C = 1.8 \text{ J/g}^\circ\text{C}$$

- 5.) To what temperature will a 50.0 g piece of glass raise if it absorbs 5275 Joules of heat and its specific heat capacity is 0.50 J/g°C? The initial temperature of the glass is 20°C.

$$5275 = (50 \text{ g}) (0.50) (F - 20)$$

$$F = 231^\circ\text{C}$$

- 6.) 25.0 grams of mercury is heated from 25°C to 155°C, and absorbs 108.75 calories of heat in the process. Calculate the specific heat of mercury in J/g°C. Assume that 4.184 Joules is equal to one calorie.

$$455 \text{ J} = (25)(C)(155-25) \quad C = 0.14 \text{ J/g}$$

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