

Specific Heat Capacity Practice Problems

Name:

1. When 3.0 kg of water is cooled from 80.0°C to 10.0°C, how much heat energy is lost?

$$c_{p,w} = Q / m \cdot \Delta T$$

$$Q = c_{p,w} (m) \Delta T$$

$$= (4186)(3)(10 - 80)$$

$$\boxed{= -879,060 \text{ J}}$$

2. How much heat is needed to raise a 0.30 kg piece of aluminum from 30.°C to 150°C?

$$Q = c_{p,a} (m_a) (\Delta T_a)$$

$$= 899 (.3) (150 - 30)$$

$$\boxed{= +32,364 \text{ J}}$$

3. Calculate the temperature change when:

- a) 10.0 kg of water loses 232 kJ of heat.

$$232 \text{ kJ} \cdot \frac{1000 \text{ J}}{1 \text{ kJ}} = 232,000 \text{ J}$$

$$-232,000 \text{ J} = (4186)(10.0 \text{ kg}) \Delta T$$

$$\frac{-232,000}{4186} = \Delta T$$

$$\boxed{8.5^\circ \text{C} = \Delta T}$$

- b) 1.96 kJ of heat are added to 500. g of copper.

$$1.96 \text{ kJ} \cdot \frac{1000 \text{ J}}{1 \text{ kJ}} = 1960 \text{ J}$$

$$500 \text{ g} \cdot \frac{1 \text{ kg}}{1000 \text{ g}} = .5 \text{ kg}$$

$$1960 \text{ J} = (387)(.5 \text{ kg}) \Delta T$$

$$\frac{1960}{193.5} = \Delta T$$

$$\boxed{10.1^\circ \text{C} = \Delta T}$$

4. $2.52 \times 10^4 \text{ J}$ of heat are added to 2.0 kg of mercury to reach a final temperature of 130°C. What was the initial temperature of the mercury?

$$T_f = 130^\circ \text{C}$$

$$Q = 2.52 \times 10^4$$

$$m = 2.0 \text{ kg}$$

$$c_{p,m} = 138$$

$$2.52 \times 10^4 = (138)(2)(130 - T_i)$$

$$2.52 \times 10^4 = 276(130 - T_i)$$

$$91.3 = 130 - T_i$$

$$-38.7 = -T_i$$

$$\boxed{38.7^\circ \text{C} = T_i}$$

Heat Change Practice Problems

Sebastian Brünink



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