

## Determining an Empirical Formula

### PRE-LAB DISCUSSION

In a sample of a compound, regardless of the size of the sample, the number of moles of one element in the sample divided by the number of moles of another element in the sample will form a small whole-number ratio. These small whole-number ratios can be used to determine the subscripts in the empirical formula of the compound.

For example, suppose that in a 24-gram sample of a compound, there are 18 g of carbon (1.5 moles) and 6 g of hydrogen (6 moles). These numbers form the small whole-number mole ratio of 4 to 1:

$$\frac{6 \text{ moles hydrogen}}{1.5 \text{ moles carbon}} = \frac{4}{1}$$

The 4-to-1 ratio means that for every 4 atoms of hydrogen in the compound, there is 1 atom of carbon. The empirical formula (the simplest whole number ratio) of the compound is  $\text{CH}_4$ .

In this experiment, the mass, in grams, of each of two elements in a compound will be experimentally determined. From this information, the empirical formula of the compound will be determined.

### PURPOSE

Demonstrate how mass relationships show that magnesium and oxygen combine in a definite whole-number ratio.

### MATERIALS and EQUIPMENT

magnesium ribbon  
ring stand  
iron ring  
clay triangle  
crucible tongs

crucible and cover  
burner  
balance  
safety goggles  
dropper pipet



### SAFETY

1)

2)

3)

### PROCEDURE

1. Measure the mass of the crucible and report in data table.
2. Cut a 20-cm length of magnesium ribbon into ~1-cm pieces. Place the pieces in the crucible and measure the mass of the crucible and its contents.
3. Cover the crucible and place it in a clay triangle like in the figure. Heat gently for 2 minutes. Using crucible tongs, carefully tilt the cover to provide an opening for air to enter the crucible. Heat the partially covered crucible strongly for 10 minutes.
4. Turn off the burner, cover the crucible, and allow the contents to cool. When the crucible is cool enough to touch, remove the cover and examine the contents.
5. Put the cover all the way on and allow to cool.
6. Add a few drops of water to the crucible.

# Empirical Formula Determination Pre Lab Question Answers

**Thirumalaisamy P.**  
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