

Electron Configuration Practice
Chemistry

Name: _____
Due Date: _____

How to write an electron configuration:

- Determine the total number of electrons to be represented.
- Use the Aufbau principle to fill the orbitals with electrons for elements 1-23. Refer to electron configuration periodic table for elements after 23
- The sum of the superscripts should equal the total number of electrons. Example:
 $12\text{Mg } 1s^2 2s^2 2p^6 3s^2$

I. Configuration Writing Practice

A. Write a ground state electron configuration for each neutral atom. Ground state means that all of the lowest possible energy levels (up to the proper number of electrons for the element) are filled.

- Na: $1s^2 2s^2 2p^6 3s^1$
- Pb: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^2$
- Sr: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2$
- U: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^3 6d^1$
- N: $1s^2 2s^2 2p^3$
- * Ag: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1 4d^{10}$
- Ti: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^2$
- Ce: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^1 5d^1$
- Cl: $1s^2 2s^2 2p^6 3s^2 3p^5$
- Hg: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10}$

Electron Configuration Quiz Answer Key

William H. Brown, Thomas Poon



Electron Configuration Quiz Answer Key:

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Table of Contents Electron Configuration Quiz Answer Key

1. Understanding the eBook Electron Configuration Quiz Answer Key
 - The Rise of Digital Reading Electron Configuration Quiz Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Quiz Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Configuration Quiz Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Quiz Answer Key
 - Personalized Recommendations
 - Electron Configuration Quiz Answer Key User Reviews and Ratings
 - Electron Configuration Quiz Answer Key and Bestseller Lists
5. Accessing Electron Configuration Quiz Answer Key Free and Paid eBooks
 - Electron Configuration Quiz Answer Key Public Domain eBooks
 - Electron Configuration Quiz Answer Key eBook Subscription Services
 - Electron Configuration Quiz Answer Key Budget-Friendly Options

6. Navigating Electron Configuration Quiz Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Electron Configuration Quiz Answer Key Compatibility with Devices
 - Electron Configuration Quiz Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electron Configuration Quiz Answer Key
 - Highlighting and Note-Taking Electron Configuration Quiz Answer Key
 - Interactive Elements Electron Configuration Quiz Answer Key
8. Staying Engaged with Electron Configuration Quiz Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Configuration Quiz Answer Key
9. Balancing eBooks and Physical Books Electron Configuration Quiz Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Configuration Quiz Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Electron Configuration Quiz Answer Key
 - Setting Reading Goals Electron Configuration Quiz Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Electron Configuration Quiz Answer Key
 - Fact-Checking eBook Content of Electron Configuration Quiz Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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