

1. Determine period and group from electron configuration below!

- $1s^2 2s^2 2p^2$
- $1s^2 2s^2 2p^4$
- $1s^2 2s^2 2p^6 3s^2 3p^2$
- $1s^2 2s^2 2p^6 3s^2 3p^4$
- $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$

2. Determine the number of atom from the period and group below!

- Group II, period 2
- Group III, period 2
- Group II, period 4
- Group IV, period 3
- Group III, period 4

3. Look the figure below!



The right statement is ...

- No. 1 is proton
- No. 2 is proton
- No. 3 is neutron
- No. 2 is electron

4. Look the figure below!



The maximum electron of $n=2$ is ...

- 2
- 8
- 18
- 32

5. Configuration of atom is $1s^2 2s^2 2p^4 3s^2 3p^4 4s^2 3d^4$.
Period and group of atom are ...

- Period 3 and group IV A
- Period 3 and group VI B
- Period 4 and group VI A
- Period 4 and group VI B

6. Atom ${}_Z X$ have shells are ...

- K, L, M, N
- K, L, M
- K, L
- K

7. Look at the picture!



The atom has charged.

- Netral
- Positif
- Negatif
- no

8. Configuration of atom is $1s^2 2s^2 2p^6 3s^2 3p^4 4s^1$. Period and group of atom are ...

- Period 3 and group I A
- Period 3 and group I B
- Period 4 and group I A
- Period 4 and group I B

9. Look the figure below!



The maximum electron of $n=3$ is

- 2
- 8
- 18
- 32

10. Look at the picture!



The atom has charged.

- Netral
- Positif
- Negatif
- no

11. The figure on the right shows the atomic shell of an atom.



The maximum numbers of electrons that can be occupied by the four shell are ...

- 2
- 8
- 18
- 32

12. Study the table on the right.

Number	The name of element
1	Chlorine
2	Aurum
3	Carbon
4	Fluorine

The elements that are not classified into halogen group are ...

- 1, 2 and 3
- 2 and 3
- 1 and 3
- 2 and 4

13. The electron configuration of an element is: 2, 8, 8, 2.

- The number of valence electron is 2
- The atom is included in alkali metal
- The period of the atom is 4
- The atom is include in earth alkali

The correct statement according to the configuration above are numbers....

- 1, 2 and 3
- 1, 2 and 4
- 1, 3 and 4
- 2, 3 and 4

14. The electron configuration of the unknown atom is 2, 8, 8, 2. The position of the atom in a periodic table is ...

- group III and period 3
- group II and period 4
- group IV and period 4
- group IV and period 3

15. The position of an atom in a periodic table is group II and period 4. The atomic number of the atom is ...

- 30
- 27
- 20
- 15

16. There are atoms belong to alkali metal group, **excepted** ...

- lithium
- sodium
- hydrogen
- potassium

Electron Configuration Test And Answer

Mark Steyvers



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