

Name \_\_\_\_\_

# Electrons in Atoms



## Study Guide: Properties of Waves and Quantum Mechanics

### Properties of Waves

- **Wavelength ( $\lambda$ ):** The distance between successive crests of a wave, usually measured in meters.
- **Frequency ( $\nu$ ):** The number of waves that pass a given point per second, measured in Hertz (Hz).
- **Amplitude:** The height of the wave crest or depth of the trough, related to the wave's energy.
- **Speed of light ( $c$ ):** The rate at which the wave propagates through a medium, calculated using  $c = \lambda\nu$

### Calculating the Wavelength of a Wave

$$\lambda = c/\nu$$

- $\lambda$  is the wavelength.
- $c$  is the wave speed.
- $\nu$  is the frequency.

### Speed of Light

- The speed of light in a vacuum ( $c$ ) is approximately  $3.00 \times 10^8$  meters per second (m/s).

### The Quantum Equation

- Relates energy of a photon to its frequency:  $E = h\nu$ 
  - $E$  is the energy.
  - $h$  is Planck's constant ( $6.626 \times 10^{-34}$  Js).
  - $\nu$  is the frequency.

### Element Emission Spectrum

The concept of quantized energy levels within atoms is crucial to understanding atomic spectra. When an electron absorbs energy, it jumps to a higher energy level, and when it returns to a lower energy level, it emits a photon of light with a specific energy corresponding to the difference in energy levels. This explains why each element has a unique spectral fingerprint, a set of specific wavelengths of light it emits or absorbs.

- When elements are heated, they emit light at specific wavelengths, creating an emission spectrum. This spectrum is unique to each element and can be used to identify them.



# Electrons In Atoms Study Guide

**Guillaume Favre**



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