

<p style="text-align: center;"><b>Rate and Growth</b></p> <p><b>Rate</b>  <math>\frac{dY}{dt}</math></p> <p><b>Population Growth</b>  <math>\frac{dN}{dt} = B - D</math></p> <p><b>Exponential Growth</b>  <math>\frac{dN}{dt} = r_{max}N</math></p> <p><b>Logistic Growth</b>  <math>\frac{dN}{dt} = r_{max}N\left(\frac{K-N}{K}\right)</math></p> <p><b>Temperature Coefficient <math>Q_{10}</math><sup>†</sup></b>  <math>Q_{10} = \left(\frac{k_2}{k_1}\right)^{\frac{10}{T_2-T_1}}</math></p> <p><b>Primary Productivity Calculation</b>  <math>\frac{\text{mg O}_2}{\text{L}} \times \frac{0.698 \text{ mL}}{\text{mg}} = \frac{\text{mL O}_2}{\text{L}}</math>  <math>\frac{\text{mL O}_2}{\text{L}} \times \frac{0.536 \text{ mg C fixed}}{\text{mL O}_2} = \frac{\text{mg C fixed}}{\text{L}}</math>            (at standard temperature and pressure)</p>	<p><math>dY</math> = amount of change  <math>dt</math> = change in time  <math>B</math> = birth rate  <math>D</math> = death rate  <math>N</math> = population size  <math>K</math> = carrying capacity  <math>r_{max}</math> = maximum per capita growth rate of population</p> <p><math>T_2</math> = higher temperature  <math>T_1</math> = lower temperature  <math>k_2</math> = reaction rate at <math>T_2</math>  <math>k_1</math> = reaction rate at <math>T_1</math>  <math>Q_{10}</math> = the factor by which the reaction rate increases when the temperature is raised by ten degrees</p>	<p><b>Water Potential (<math>\Psi</math>)</b>  <math>\Psi = \Psi_p + \Psi_s</math></p> <p><math>\Psi_p</math> = pressure potential  <math>\Psi_s</math> = solute potential</p> <p>The water potential will be equal to the solute potential of a solution in an open container because the pressure potential of the solution in an open container is zero.</p> <p><b>The Solute Potential of a Solution</b>  <math>\Psi_s = -iCRT</math></p> <p><math>i</math> = ionization constant (this is 1.0 for sucrose because sucrose does not ionize in water)  <math>C</math> = molar concentration  <math>R</math> = pressure constant (<math>R = 0.0831</math> liter bars/mole K)  <math>T</math> = temperature in Kelvin (<math>^{\circ}\text{C} + 273</math>)</p>
<p style="text-align: center;"><b>Surface Area and Volume</b></p> <p><b>Volume of a Sphere</b>  <math>V = \frac{4}{3}\pi r^3</math></p> <p><b>Volume of a Rectangular Solid</b>  <math>V = lwh</math></p> <p><b>Volume of a Right Cylinder</b>  <math>V = \pi r^2 h</math></p> <p><b>Surface Area of a Sphere</b>  <math>A = 4\pi r^2</math></p> <p><b>Surface Area of a Cube</b>  <math>A = 6s^2</math></p> <p><b>Surface Area of a Rectangular Solid</b>  <math>A = \sum</math> surface area of each side</p>	<p><math>r</math> = radius  <math>l</math> = length  <math>h</math> = height  <math>w</math> = width  <math>s</math> = length of one side of a cube  <math>A</math> = surface area  <math>V</math> = volume  <math>\Sigma</math> = sum of all</p>	<p><b>Dilution (used to create a dilute solution from a concentrated stock solution)</b>  <math>C_i V_i = C_f V_f</math></p> <p><math>i</math> = initial (starting)      <math>C</math> = concentration of solute  <math>f</math> = final (desired)      <math>V</math> = volume of solution</p> <p><b>Gibbs Free Energy</b>  <math>\Delta G = \Delta H - T\Delta S</math>  <math>\Delta G</math> = change in Gibbs free energy  <math>\Delta S</math> = change in entropy  <math>\Delta H</math> = change in enthalpy  <math>T</math> = absolute temperature (in Kelvin)</p> <p><math>\text{pH}^* = -\log_{10} [\text{H}^*]</math></p>
<p>* For the purposes of the AP Exam, students will not be required to perform calculations using this equation; however, they must understand the underlying concepts and applications.</p> <p>† For use with labs only (optional).</p>		

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**Thirumalaisamy P.  
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