

**Answer Key****Naming and Writing Formulas for Ionic Compounds**

Write the names for the following compounds:

1.  $\text{KClO}_4$  – potassium perchlorate
2.  $\text{Sr}_3\text{P}_2$  – strontium phosphide
3.  $\text{NH}_4\text{NO}_2$  – ammonium nitrite
4.  $\text{NaCl}$  – sodium chloride
5.  $\text{CoS}$  – cobalt (II) sulfide
6.  $\text{ZnCl}_2$  – zinc (II) chloride
7.  $\text{Li}_3\text{N}$  – lithium nitride
8.  $\text{BaSO}_3$  – barium sulfite
9.  $\text{Co}_2\text{S}_3$  – cobalt (III) sulfide
10.  $\text{CsF}$  – cesium fluoride
11.  $\text{NaNO}_3$  – sodium nitrate
12.  $\text{CuI}_2$  – copper (II) iodide
13.  $\text{TiO}_2$  – titanium (IV) oxide
14.  $\text{KMnO}_4$  – potassium permanganate
15.  $\text{ZnCN}_2$  – zinc (II) cyanide
16.  $\text{Al}_2(\text{CO}_3)_3$  – aluminum carbonate
17.  $\text{Fe}_2(\text{CO}_3)_3$  – iron (III) carbonate
18.  $\text{BaCrO}_4$  – barium chromate
19.  $\text{Pb}_3(\text{PO}_4)_2$  – lead (II) phosphate
20.  $\text{NaHCO}_3$  – sodium hydrogen carbonate or sodium bicarbonate

Write the formulas for the following compounds:

1. cesium bromide –  $\text{CsBr}$
2. aluminum chlorite –  $\text{Al}(\text{ClO}_2)_3$
3. ammonium phosphite –  $(\text{NH}_4)_3\text{PO}_3$
4. copper (II) nitrate –  $\text{Cu}(\text{NO}_3)_2$
5. sodium fluoride –  $\text{NaF}$
6. lead (IV) sulfate –  $\text{Pb}(\text{SO}_4)_2$
7. calcium carbonate –  $\text{CaCO}_3$
8. barium oxide –  $\text{BaO}$
9. iron (III) sulfide –  $\text{Fe}_2\text{S}_3$
10. manganese (II) chromate –  $\text{MnCrO}_4$

# Formula Of Ionic Compound Lab Answer Key

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