

EXERCISES

MULTIPLE-CHOICE QUESTIONS

8) Consider the following reaction:



Which one of the following statements is correct?

- A. Since very large, the reaction is irreversible.
- B. Changing the pressure will affect the value of K_c .
- C. In an open beaker, the reaction is effectively irreversible.
- D. Changing the temperature will not affect the value of K_c .

9) The following reaction, in which nitrogen monoxide, NO , is converted to nitrogen dioxide, NO_2 , occurs at equilibrium:



In a given closed vessel, the equilibrium constant for this reaction is 0.250. If the initial molar concentrations of NO and NO_2 at equilibrium are 0.2000 M and three times the initial concentration of NO_2 at equilibrium is

- A. 0.0500 M
- B. 0.0600 M
- C. 0.0400 M
- D. 0.0300 M

10) Consider carbon dioxide, CO_2 , in contact with water, H_2O , in a closed, constant-volume, H_2CO_3 , H_2CO_3 , equilibrium vessel with H_2O and the concentration of H_2CO_3 and hydrochloric acid, HCl , according to the equilibrium below:



At equilibrium the concentration of carbonic acid is 0.0010 M and the concentration of H^+ ions is 0.0010 M. If the concentration of CO_2 is 0.0300 M, the concentration of H_2CO_3 is 0.0010 M. If the concentration of H^+ ions is 0.0010 M, the concentration of HCO_3^- ions is 0.0010 M. Which of the following statements is most likely to result in an increase in the concentration of H^+ ions?

- A. Change the volume.
- B. Double the mass CO_2 gas.
- C. Add a few drops of perchloric acid solution.
- D. Add a few drops of 0.1 M sodium hydroxide, NaOH , solution.

Equilibrium Test 2 Multiple Choice Answers

Ronald Pirayoff



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