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An Introduction to Management Science

Quantitative Approaches to Decision Making



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 $2 \text{C}_6\text{H}_{10} + 17 \text{O}_2 \rightarrow 12 \text{CO}_2 + 10 \text{H}_2\text{O}$
 $3 \text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $4 \text{Na}_2\text{SO}_3 + \text{O}_2 \rightarrow 4 \text{Na}_2\text{SO}_4$
 $5 \text{NF}_3 + 2 \text{N}_2\text{O}_3 \rightarrow 5 \text{N}_2 + 3 \text{O}_2$
 using the equation from problem 2 above answer the following

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 $3 \text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $2 \text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$
 using the following equation

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 $\text{Ce} + \text{Zn} + \text{CuCl}_2 \rightarrow \text{ZnCl}_2 + \text{Cu} + \text{Zn} + \text{CuCl}_2$
 $2 \text{Zn} + \text{Cu} \rightarrow 2 \text{ZnCl}_2 + \text{Cu}$
 how many moles of $\text{Ce} + \text{ZnCl}_2 + \text{ZnCl}_2$ will be produced from 23.0 g of $\text{Ce} + \text{Zn}$ assuming $\text{Ce} + \text{CuCl}_2 + \text{CuCl}_2$ is available in excess

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 $2 \text{C}_6\text{H}_{10} + 17 \text{O}_2 \rightarrow 12 \text{CO}_2 + 10 \text{H}_2\text{O}$
 $3 \text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $4 \text{Na}_2\text{SO}_3 + \text{O}_2 \rightarrow 4 \text{Na}_2\text{SO}_4$
 $5 \text{NF}_3 + 2 \text{N}_2\text{O}_3 \rightarrow 5 \text{N}_2 + 3 \text{O}_2$
 using the following equation
 $2 \text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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