

Gas Laws HOMEWORK

Lesson 1 Homework

1. Define pressure. *Force that a gas exerts on its container*

2. Convert the following:

a. 100mmHg to atm

b. 100mmHg to kPa

c. 1.5atm to torr

d. 2.3×10^5 Pa to atm

e. 20kPa to atm

101) $100 \text{ mmHg} \times \frac{1 \text{ atm}}{760 \text{ mmHg}} = 0.132 \text{ atm}$

102) $100 \text{ mmHg} \times \frac{1.01325 \text{ kPa}}{760 \text{ mmHg}} = 0.133 \text{ kPa}$

103) $2.3 \times 10^5 \text{ Pa} \times \frac{1 \text{ atm}}{101325 \text{ Pa}} = 2.27 \text{ atm}$

104) $1.5 \text{ atm} \times \frac{760 \text{ torr}}{1 \text{ atm}} = 1140 \text{ torr}$

105) $20 \text{ kPa} \times \frac{1 \text{ atm}}{101.325 \text{ kPa}} = 0.197 \text{ atm}$

3. Calculate the atmospheric pressure in kPa on a day when the pressure is 760mm.

106) $760 \text{ mm} \times \frac{101.325 \text{ kPa}}{760 \text{ mmHg}} = 101.325 \text{ kPa}$

4. Calculate the pressure of a sample of gas (in mmHg) inside an inverted manometer if the mercury column is 30mm higher on the side of the closed tube than on the atmosphere. The atmospheric pressure is 760mm. Draw a picture to illustrate your answer.

107) $760 \text{ mm} + 30 \text{ mm} = 790 \text{ mmHg}$

107) The mercury is 30mm HIGHER than the atmospheric pressure, then it's 30mmHg HIGHER than 760 mmHg.

108) The pressure should decrease a centimeter with the height of the mercury. Then when the atmospheric level of 764.5mm.

108) $760 \text{ mmHg} - 30 = 730 \text{ mmHg}$

5. If a gas at a pressure of 100mm in a 1.0L container has its pressure increased to 400mm, what is the new volume of the container? (Assume temperature is constant).

$P_1 V_1 = P_2 V_2$

$(100 \text{ mm})(1.0 \text{ L}) = (400 \text{ mm})(V_2)$

$V_2 = 0.25 \text{ L}$

6. What is the new pressure if a 5.00 closed container with a pressure of 2.0atm is increased to 0.50L?

$P_1 V_1 = P_2 V_2$

$(2 \text{ atm})(5 \text{ L}) = P_2(0.5 \text{ L})$

$P_2 = 20 \text{ atm}$

7. When 0.5 L of nitrogen and 0.5 L of hydrogen both at 0°C and 1.0 atm, are pumped into an evacuated 1.0 L rigid container, what is the final pressure in the container at 0°C? *Final temperature doesn't change, you use both of part of the equation.*

For final pressure of Nitrogen:

$P_1 V_1 = P_2 V_2$

$(1 \text{ atm})(0.5 \text{ L}) = P_2(1 \text{ L})$

$P_2 = 0.5 \text{ atm}$

For final pressure of Hydrogen:

$P_1 V_1 = P_2 V_2$

$(1 \text{ atm})(0.5 \text{ L}) = P_2(1 \text{ L})$

$P_2 = 0.5 \text{ atm}$

Total Pressure = $P_1 + P_2 + P_3$

$1.0 = 0.5 + 0.5 = 1.0 \text{ atm}$

Atmospheric pressure: 101325 Pa (1.01325 atm)

$P_1 V_1 = P_2 V_2$

$(1 \text{ atm})(0.5 \text{ L}) = P_2(1 \text{ L})$

$P_2 = 0.5 \text{ atm}$

Final total pressure: 1.01325 atm + 0.5 atm = 1.51325 atm

8. A sealed flask contains a gas with pressure P. If the number of moles of the gas is doubled, the absolute temperature is doubled and the volume of the container is quadrupled, what is the new pressure in terms of P?

Doubling the moles doubles the pressure.

$P_1 V_1 = P_2 V_2$

Doubling the volume cuts the pressure in 1/2.

$P_1 V_1 = P_2 V_2$

quadrupling volume cuts the pressure in 1/4.

$P_1 V_1 = P_2 V_2$

9. If 4 moles of gas are added to a container that already holds 2 moles of gas, how will the pressure change within the container?

$P_1 V_1 = P_2 V_2$

$P_1 V_1 = P_2 V_2$

$P_2 = 2P_1$

Adding three 2 moles to 2 moles results in an increase of pressure by 2 times.

10. If a balloon containing 100L of gas at 40 °C and 100 mmHg rises to an altitude where the pressure is 50 mmHg and the temperature is 20 °C, what is the new volume of the balloon (in L)?

$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$

$\frac{(100 \text{ mmHg})(100 \text{ L})}{(40 \text{ }^\circ\text{C})} = \frac{(50 \text{ mmHg})(V_2)}{(20 \text{ }^\circ\text{C})}$

$V_2 = 100 \text{ L}$

11. A sample of nitrogen gas has a pressure of 87.5 mmHg in a 500. mL flask. What is the pressure of this gas sample when it is transferred to a 225.0 mL flask at the same temperature?

$P_1 V_1 = P_2 V_2$

$(87.5 \text{ mmHg})(500 \text{ mL}) = P_2(225 \text{ mL})$

$P_2 = 192 \text{ mmHg}$

12. You have 4.5 L of NO at a temperature of 22 °C. What volume would the NO occupy at 47 °C? (Assume the pressure is constant).

$\frac{V_1}{T_1} = \frac{V_2}{T_2}$

$\frac{4.5 \text{ L}(22 \text{ }^\circ\text{C} + 273)}{22 \text{ }^\circ\text{C}} = \frac{V_2(47 \text{ }^\circ\text{C} + 273)}{47 \text{ }^\circ\text{C}}$

$V_2 = 6.88 \text{ L}$

13. You have a sample of a gas in a flask with a volume of 240 mL. At 20 °C the pressure of the gas is 100 mmHg. If you decrease the temperature to -40 °C, what is the gas pressure at the lower temperature? *Final volume doesn't change, you use both of part of the equation.*

$P_1 V_1 = P_2 V_2$

$(100 \text{ mmHg})(240 \text{ mL}) = P_2(240 \text{ mL})$

$P_2 = 400 \text{ mmHg}$

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