

## Gas Laws Packet #2

### Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, " $PV = nRT$ ", and the universal gas constant  $R = 0.0821 \frac{L \cdot atm}{K \cdot mol}$  to solve the following problems:

If pressure is needed in kPa then convert by multiplying by  $101.3 kPa / 1 atm$  to get  
 $R = 8.31 \frac{L \cdot kPa}{K \cdot mole}$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

**204.6 K**

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of 37 °C, how many moles of gas do I have?

**1.26 mol**

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

**1.64 atm**

- 4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of 56 °C, what is the volume of the container that the gas is in?

**2311 L**

- 5) If I have 17 moles of gas at a temperature of 67 °C, and a volume of 33.39 liters, what is the pressure of the gas?

**5.4 atm or 540 kPa**

- 6) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

**.5 mol**

# Gas Laws Review Packet Answers

**K Morrison**



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