

Gas Law Stoichiometry Worksheet

Directions: Use significant figures and units in the problems below.

$PV = nRT$

SHOW ALL WORK!!!

1. Given the following unbalanced chemical equation for the combination reaction of sodium metal and chlorine gas:



- a. What volume of chlorine gas, measured at STP, is necessary for the complete reaction of 4.81 g of sodium metal.

$4.81 \text{ g Na} \left| \frac{1 \text{ mol}}{23.0 \text{ g Na}} \right| \left| \frac{1 \text{ mol Cl}_2}{2 \text{ mol Na}} \right| \left| \frac{22.4 \text{ L}}{1 \text{ mol}} \right| = 2.34 \text{ L Cl}_2$

$2.34 \text{ L Cl}_2$



- a. Balance the above reaction.  
 b. What volume of oxygen gas at 25°C and 1.04 atm is needed for the complete combustion of 5.53 g of propane?

$n = x$   
 $V = O_2(x)$

$T_1 = 25 + 273 = 298 \text{ K}$   
 $P_1 = 1.05 \text{ atm}$

$5.53 \text{ g C}_3\text{H}_8 \left| \frac{1 \text{ mol}}{44 \text{ g}} \right| \left| \frac{5 \text{ mol O}_2}{1 \text{ mol C}_3\text{H}_8} \right| = 0.628 \text{ moles}$   
 $V = 14.6 \text{ L}$

$14.6 \text{ L}$

3. Potassium permanganate,  $\text{KMnO}_4$ , is produced commercially by oxidizing aqueous potassium manganate,  $\text{K}_2\text{MnO}_4$ .



- a. Balance the above reaction.  
 b. What volume of  $\text{Cl}_2(g)$ , measured at STP, is needed to produce 10.0 g of  $\text{KMnO}_4$ ?

$V = \frac{nRT}{P}$

$10.0 \text{ g KMnO}_4 \left| \frac{1 \text{ mol}}{158 \text{ g}} \right| \left| \frac{1 \text{ mol Cl}_2}{2 \text{ mol}} \right| \left| \frac{22.4 \text{ L}}{1 \text{ mol Cl}_2} \right| = 0.709 \text{ L Cl}_2$

$0.709 \text{ L}$

4. If water is added to magnesium nitride, ammonia gas is produced when the mixture is heated.



- a. Balance the above reaction.  
 b. If 10.3 g of magnesium nitride is treated with 10.3 g of water, what volume of ammonia gas would be collected at 24°C and 752 mmHg?

$10.3 \text{ g Mg}_3\text{N}_2 \left| \frac{1 \text{ mol}}{100.9 \text{ g}} \right| \left| \frac{2 \text{ mol NH}_3}{1 \text{ mol}} \right| = 0.204 \text{ mol}$   
 Lim Reagent

$10.3 \text{ g H}_2\text{O} \left| \frac{1 \text{ mol}}{18 \text{ g}} \right| \left| \frac{2 \text{ mol NH}_3}{3 \text{ mol}} \right| = 0.381 \text{ mol}$

$24 + 273 = 297 \text{ K}$   
 $752 / 760 = 0.989 \text{ atm}$

$5.03 \text{ L}$

$V = \frac{nRT}{P}$

$V = \frac{(0.204 \text{ mol} + 0.082 \text{ L} \cdot \text{atm} / 297 \text{ K})}{0.989 \text{ atm}} = 5.03 \text{ L}$

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**Yijin Wang**



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