

KEY

Gas Stoichiometry Worksheet

Directions: Use the gas laws we have learned to solve each of the following problems. Each of the chemical equations must first be balanced. Show all your work for credit.

1. When calcium carbonate is heated strongly, carbon dioxide gas is released according to the following equation:



What volume of $\text{CO}_2(g)$, measured at STP, is produced if 15.2 grams of $\text{CaCO}_3(s)$ is heated?

$$\frac{15.2 \text{ g CaCO}_3}{100.1 \text{ g}} \times \frac{1 \text{ mol CaCO}_3}{1 \text{ mol CaCO}_3} \times \frac{1 \text{ mol CO}_2}{1 \text{ mol CaCO}_3} \times \frac{22.4 \text{ L CO}_2}{1 \text{ mol CO}_2} = 3.40 \text{ L CO}_2$$

2. The synthesis of sodium chloride occurs according to the reaction:



What volume of chlorine at STP is necessary for the complete reaction of 4.81 grams of sodium metal?

$$\frac{4.81 \text{ g Na}}{23.0 \text{ g}} \times \frac{1 \text{ mol Na}}{2 \text{ mol Na}} \times \frac{1 \text{ mol Cl}_2}{2 \text{ mol Na}} \times \frac{22.4 \text{ L Cl}_2}{1 \text{ mol Cl}_2} = 2.34 \text{ L Cl}_2$$

3. Potassium permanganate is produced commercially by the this reaction.



What volume of chlorine gas at STP would be required to produce 10.0 grams of KMnO_4 ?

$$\frac{10.0 \text{ g}}{158.0 \text{ g}} \times \frac{1 \text{ mol KMnO}_4}{2 \text{ mol KMnO}_4} \times \frac{1 \text{ mol Cl}_2}{2 \text{ mol KMnO}_4} \times \frac{22.4 \text{ L}}{1 \text{ mol}} = 7.09 \text{ liters}$$

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Table of Contents Gas Stoichiometry Problems With Answers

1. Understanding the eBook Gas Stoichiometry Problems With Answers
 - The Rise of Digital Reading Gas Stoichiometry Problems With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Stoichiometry Problems With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Stoichiometry Problems With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Stoichiometry Problems With Answers
 - Personalized Recommendations
 - Gas Stoichiometry Problems With Answers User Reviews and Ratings
 - Gas Stoichiometry Problems With Answers and Bestseller Lists
5. Accessing Gas Stoichiometry Problems With Answers Free and Paid eBooks
 - Gas Stoichiometry Problems With Answers Public Domain eBooks
 - Gas Stoichiometry Problems With Answers eBook Subscription Services
 - Gas Stoichiometry Problems With Answers Budget-Friendly Options
6. Navigating Gas Stoichiometry Problems With Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Gas Stoichiometry Problems With Answers Compatibility with Devices
 - Gas Stoichiometry Problems With Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Stoichiometry Problems With Answers
 - Highlighting and Note-Taking Gas Stoichiometry Problems With Answers
 - Interactive Elements Gas Stoichiometry Problems With Answers
 8. Staying Engaged with Gas Stoichiometry Problems With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Stoichiometry Problems With Answers
 9. Balancing eBooks and Physical Books Gas Stoichiometry Problems With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Stoichiometry Problems With Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Gas Stoichiometry Problems With Answers
 - Setting Reading Goals Gas Stoichiometry Problems With Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Gas Stoichiometry Problems With Answers
 - Fact-Checking eBook Content of Gas Stoichiometry Problems With Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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