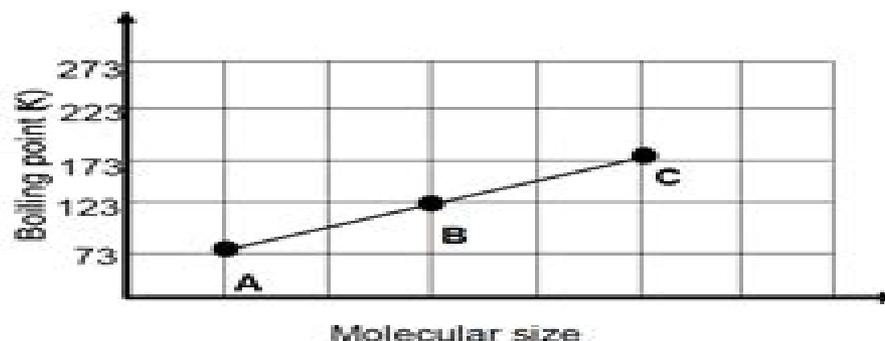


**QUESTION 2**

2.1 The graph of molecular size versus the boiling point is given below.

The letters **A**, **B** and **C** represent the compounds  $\text{CH}_4$ ,  $\text{C}_2\text{H}_6$  and  $\text{C}_3\text{H}_8$  respectively.



2.1.1 Define the term *boiling point*. (2)

2.1.2 Describe the trend in the boiling points of the compounds as shown by the graph. (2)

2.1.3 Explain the answer to QUESTION 2.1.2 by referring to MOLECULAR SIZE, TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)

2.1.4 Which ONE of the compounds (**A**, **B** or **C**) has the HIGHEST vapour pressure? Explain the answer by referring to the data on the graph. (2)

2.2 Consider the two molecules in the table below.

| NAME OF SUBSTANCE | FORMULA       | MOLACULAR MASS (g) | BOILING POINT ( $^{\circ}\text{C}$ ) |
|-------------------|---------------|--------------------|--------------------------------------|
| Ammonia           | $\text{NH}_3$ | 17                 | - 33                                 |
| Phosphine         | $\text{PH}_3$ | x                  | - 87,4                               |

2.2.1 Determine the molecular mass (x) of phosphine (1)

2.2.2 Explain the difference in the boiling points by referring to the TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)

[13]

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