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# How To Make A Manual In Word

**Chao Zhang**



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web as a result of the reaction the equilibrium concentration of  $NH_3(g)$  became  $0.540\text{ M}$  what is the value of the equilibrium constant for this reaction at the given temperature of  $700\text{ C}$   $N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$  i write the initial concentrations of reactants and products

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web equilibrium 1  $864\text{ M}$  1  $364\text{ M}$  0  $636\text{ M}$  equilibrium 1  $600\text{ M}$  1  $000\text{ M}$  0  $400\text{ M}$  trial 5 a  $g$   $b$   $g$   $c$   $g$  trial 6 a  $g$   $b$   $g$   $c$   $g$  initial 1  $400\text{ M}$  1  $200\text{ M}$  0  $400\text{ M}$  initial 0  $750\text{ M}$  2  $000\text{ M}$  0  $250\text{ M}$  change change equilibrium 1  $388\text{ M}$  1  $188\text{ M}$  0  $412\text{ M}$  equilibrium 0  $675\text{ M}$  1  $925\text{ M}$  0  $325\text{ M}$  1 examine model 1 a

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