

GENERAL CHEMISTRY 2  
CHAPTER TEST 2

Name: \_\_\_\_\_  
Section: \_\_\_\_\_

Score: \_\_\_\_\_  
Date: \_\_\_\_\_

Direction: Choose the letter of the correct answer. (*Strictly no erasures on your answer sheets.*)

- It is a special type of homogeneous mixture composed of two or more substances.  
a. suspension      b. **solution**      c. colloids      d. alloys
- Which of the following is an example of liquid in a solid solution?  
a. Air      b. Brass      c. vinegar      d. **dental amalgam**
- You are tasked to analyze a water samples containing contaminants. In what unit of concentration you are going to report the results of analysis?  
a. **parts per million**      b. percent      c. Molarity      d. molality
- What is the mole fraction of solvent in 40% isopropyl alcohol solution? (Molar mass isopropyl = 60.1 g/mol; molar mass of water = 18.0 g/mole)  
a. 0.16      b. 0.67      c. **0.84**      d. 3.3
- What is the concentration, in ppm, if 0.025 g KCl is dissolved in 100 grams of water?  
a.  $2.5 \times 10^{-4}$  ppm      b. 2.5 ppm      c.  $4.0 \times 10^2$  ppm      d. **250 ppm**
- Which of the following statement is TRUE about molality?  
a. Molality is temperature dependent  
b. Molality is represented by moles / L  
c. Molality gives a less accurate value than molarity.  
d. **Molality gives a more accurate value than molarity.**
- The term homogeneous mixtures signify that  
a. its properties are uniform throughout the mixture.  
b. its composition is uniform throughout the mixture.  
c. **Both composition and properties are uniform throughout the mixture.**  
d. neither composition nor properties are uniform throughout the mixture.
- Which of the following is a quantitative description of the solution?  
a. Diluted      b. concentrated      c. saturated      d. **molality**
- Calculate the molarity of liquid HCl will be if density of the solution is 1.17 gm/cc. (MM of HCl= 36.46 g/mol; 1cc = 1 mL, 1 L= 1000 mL)  
a. 31.16 M      b. **32.09 M**      c. 31.16 m      d. 32.09 m
- What is the molarity of 20% (w/v) sulfuric acid ( $H_2SO_4$ ) solution?  
a. 0.51 M      b. 1.02 M      c. **2.04 M**      d. 4.08 M
- Which of the following is used as an indicator in the titration of a weak acid and a strong base?  
a. Methyl orange (3 to 4.6)      c. Methyl red (5 to 6.9)  
b. Bromothymol blue (6 to 7.5)      d. **Phenolphthalein (8 to 9.6)**
- Arrange the procedure in based on the correct sequence of the titration process.  
I. Indicator is added to the solution.  
II. A specified amount of analyte is placed on the Erlenmeyer flask.  
III. Titrate the sample until a distinct change in the color of the solution is observed.  
IV. Titrant is placed in a burette to titrate the sample containing indicator.  
a. I, II, III and IV      b. II, III, IV and I      c. III, IV, II and I      d. **II, I, IV and III**
- The following is an indication that the end-point of titration is reached except  
a. **The solution becomes dark pink in color.**  
b. All the acid has been neutralized by the base.  
c. A distinct change in the color of an indicator is noticeable.  
d. The solution changes in color from clear to light pink solution.
- What is the concentration of the sulphuric acid solution if 50 ml of 0.5 M  $Ba(OH)_2$  solution is added to 100 ml of the solution?  
a. **0.25 M**      b. 0.5 M      c. 25 M      d. 50 m
- Find the concentration of HCl, if 10 mL of 0.5 M  $Ca(OH)_2$  is required to titrate 50 mL of HCl.  
 $2HCl(aq) + Ca(OH)_2(aq) \rightarrow CaCl_2(aq) + 2H_2O(l)$   
a. 0.1 M      b. **0.2 M**      c. 0.5 M      d. 1.0 M
- Using the following equation:  
 $2NaOH + H_2SO_4 \rightarrow 2H_2O + Na_2SO_4$   
How many grams of sodium sulfate will be formed if you start with 1.25 L of a 4.0 M sodium hydroxide? (Atomic mass: Na=23.0, O=16.0, H=1.01, S=32.0.)  
a. 1420 g      b. 710.0 g      c. **155.0 g**      d. 0.7040 g

# Exam Chapter 2 nuclear Chemistry

**Sabine Zange**



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