

Study Guide: Chemical Formulas & Reactions

Key Concepts

- **Chemical Formulas:** Symbolic representations of the composition of a substance.
 - **Subscripts:** Indicate the number of atoms of each element in a molecule.
 - **Coefficients:** Indicate the relative number of moles of a substance in a reaction.
- **Chemical Reactions:** Processes that lead to the transformation of one set of chemical substances to another.
 - **Reactants:** Substances that start a chemical reaction.
 - **Products:** Substances that are produced in a chemical reaction.
- **Chemical Equations:** Symbolic representations of chemical reactions.
 - **Balanced equations:** Have the same number of atoms of each element on both sides of the equation.
- **Types of Chemical Reactions:**
 - **Synthesis:** $A + B \rightarrow AB$
 - **Decomposition:** $AB \rightarrow A + B$
 - **Single-displacement:** $A + BC \rightarrow AC + B$
 - **Double-displacement:** $AB + CD \rightarrow AD + CB$
 - **Combustion:** Reaction with oxygen,² often producing heat and light.
- **Evidence of Chemical Reactions:**
 - Change in temperature (exothermic or endothermic)
 - Formation of a precipitate
 - Production of gas
 - Change in color
 - Change in odor
- **Law of Conservation of Mass:** Matter cannot be created or destroyed in a chemical reaction.
- **Catalyst:** A substance that speeds up a reaction without being consumed.
- **Activation Energy:** The minimum amount of energy required to start a chemical reaction.

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