

EXTRA PRACTICE PROBLEMS FOR MODULE #6

1. Nitric acid is manufactured with the following reaction:



How many grams of nitrogen dioxide are required to produce 5.89×10^3 kg HNO_3 in excess water?

2. Tungsten metal, W, is produced by the following reaction:



How many grams of tungsten can be made from 1.0×10^6 g of hydrogen in excess WO_3 ?

3. How many grams of WO_3 would be used in the above reaction under the conditions specified in #2?
 4. Carbon tetrachloride is made according to the following reaction:



How many grams of CS_2 are needed to make 16 kg of CCl_4 ?

5. To make a phosphorus fertilizer ($\text{CaH}_4\text{P}_2\text{O}_8$), agricultural companies use the following reaction:



If 2.50×10^3 grams of $\text{Ca}_3\text{P}_2\text{O}_8$ are reacted with excess H_2SO_4 and H_2O , how many grams of fertilizer can be made?

6. In the following reaction:



What mass of Mg must you use to make 5.6×10^4 g of Mg(OH)_2 ?

7. Which of the following are empirical formulas?

a. SO_3

b. $\text{K}_2\text{S}_2\text{O}_3$

c. $\text{C}_6\text{H}_{12}\text{O}_6$

d. P_2H_4

8. Decomposition of 100.0 g of a dark green powder yields 39.7 g K, 27.9 g Mn, and an unknown amount of oxygen. What is the empirical formula?

9. The mass of the molecule above is 197.1 amu. What is the molecular formula?

10. A compound has an empirical formula of CH_2O . If its molecular mass is 60.0 amu, what is its molecular formula?

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Robert R. Redfield



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Aerospace Medicine and Biology, 1991. A selection of annotated references to unclassified reports and journal articles that were introduced into the NASA scientific and technical information system and announced in Scientific and Technical Aerospace Reports (STAR) and International Aerospace Abstracts (IAA).

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