

Exploring the Periodic Table

The Periodic Table was first published in 1869 by a Russian scientist named Dmitri Mendeleev. There were about 55 known elements at the time, and Mendeleev began to notice patterns in their properties when he arranged them a certain way. Originally, Mendeleev arranged the elements in order of increasing atomic mass, but scientists in the 1900s realized that it was more accurate to arrange them in order of increasing atomic number. An element's atomic number is based on how many protons it has in its nucleus. Neutral atoms have the exact same number of electrons as they do protons. Each square on the periodic table tells us certain things: the element's name, its one or two letter chemical symbol, the atomic number, and the atomic mass. The atomic mass is the average mass of one atom of the element. It can also be used to help us determine how many neutrons are in each atom of that element if we use the following formula: $\text{atomic mass} - \text{atomic \#} = \text{number of neutrons}$. Each row on the periodic table is called a period. Each column is called a group or a family. And just like members in a real family, each group shares certain characteristics. The periodic table is also divided into three main groups: metals, nonmetals, and metalloids, which has properties of both metals and nonmetals. Today's periodic table has 118 named elements.



True or False

If the answer is true, write "true" on the line. If the answer is false, replace the underlined word or phrase with one that will make the sentence correct. Write the new word(s) on the line.

- _____ The modern periodic table organizes elements based on their atomic mass.
- _____ Elements in the same period have similar chemical properties.

Exploring The Periodic Table Worksheet

JL Elias



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