

**Stoichiometry Limiting Reagent Answer Key**

You need to teach yourself about limiting reagents so that you can answer the following questions. These are the types of questions that will be on the test. You will need to use all you have learned so far and then learn the new concept of limiting reagents. In teaching yourself limiting reagents I have given you a one-two-punch approach to learning this topic on your own. The first is a series of guided questions and second is a list of resources. Together with both of these resources and your own intelligences and ingenuity, you can learn limiting reagents. Be sure to use all your resources.

The resources are as follows:

1. Your chemistry textbook chapter 12.3 pages 368-371
2. Your chemistry textbook web site.
3. Searching the internet. There are several excellent sites. I have found many by searching google and youtube for "limiting reagent tutorial" and choosing a site and learn what it teaches. Sample Sites include the following (I suggest visiting them in this order):
  - a. <http://www.wiley.com/college/chem/spencer053872/tutorial/limreagent/limreagent.html>
  - b. <http://www.iun.edu/~cpsnhd/C101webnotes/gaschem/limitingreagent.html>
  - c. <https://www.youtube.com/watch?v=j2O4DUboqDg>

The guided questions are as follows:

1. Answer the following questions using your resources.
  - a. What is a limiting reagent (reactant)? The substance/reactant that gets used up first
  - b. Is the following sentence true or **false**? A chemical reaction stops before the limiting reagent is used up. false
  - c. Circle the letter of the term that correctly completes the sentence. The reactant that is not completely used up in a chemical reaction is called the \_\_\_\_\_.
    - a. spectator reagent b. **excess reagent** c. limiting reagent d. catalyst
2. What is the significance of the limiting reagent in a reaction? It dictates the amount (moles, grams, or liters) of product that is produced
3. What happens to the amount of any reagent that is present in an excess? There is an excess (unused) amount of this reactant
4. How would you identify a limiting reagent in a chemical reaction? Calculate the # of moles of one of the products from each reactant. The reactant that produces the smallest # of moles of product is the limiting reactant/reagent.
5. In a reaction chamber, 3.0 mol of aluminum is allowed to react with 5.3 mol Cl<sub>2</sub> and reacts. The reaction is described by the following balanced chemical equation.



- a. Identify the limiting reagent for the reaction. Show your work

Al

# Explore Learning Limiting Reactants Actual Answer Key

**Marco Cascella**



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