

# Equation Writing and Predicting Products

## Chemistry I Acc

### Introduction:

If you examine your bicycle after it has been left out in the rain a number of times you will find that it has begun to rust. Rust is a slow chemical reaction of the iron in the bicycle with oxygen. If a piece of sodium is put in water, a much more rapid chemical reaction occurs. Sodium reacts with water to produce sodium hydroxide and hydrogen gas. During this reaction, enough heat is liberated to ignite the hydrogen gas, causing it to explode.

Chemists observe what is happening in a chemical reaction and try to describe it in language that is clear and simple. A chemical equation uses formulas and symbols to describe the substances involved in a reaction, the physical state of the substance, the use of a catalyst, and relative proportions. The general form of an equation is:



In this investigation, you will perform a series of reactions and make careful observations of the changes that occur. Using simple tests and your knowledge of chemistry, you will determine the identity of the products. With this information, you will write and balance chemical equations to describe the reactions.

### Materials

goggles  
magnesium ribbon  
triangular file  
copper strip  
cobalt chloride paper

lab burner  
watch glass  
glycerin  
bent glass tube  
test tube tongs

crucible tongs  
6 test tubes  
wooden splints  
rubber tubing  
graduated cylinder

manganese(IV) oxide ( $\text{MnO}_2$ )  
lead(II) nitrate ( $\text{Pb}(\text{NO}_3)_2$ ) 0.1M  
limewater ( $\text{Ca}(\text{OH})_2$  solution)  
ammonium carbonate ( $(\text{NH}_4)_2\text{CO}_3$ )

potassium iodide (KI) 0.1M  
copper(II) carbonate ( $\text{CuCO}_3$ )  
one-hole rubber stopper  
hydrogen peroxide ( $\text{H}_2\text{O}_2$ ) 10% solution

### Safety Warnings

Wear your goggles at all times during this investigation. Avoid looking directly at the burning Magnesium. The bright light could seriously damage your eyes. Tie back loose hair. Hydrochloric acid is corrosive. Avoid direct contact with it. If you get some on your skin, immediately wash it off with copious amounts of water. Ammonia is a skin and respiratory irritant so avoid inhaling it deeply. Lead compounds are poisonous so be sure to avoid contact with skin. If contact occurs immediately wash it off with copious amounts of water and inform your teacher. Glass tubing breaks easily. Exercise caution when working with it and use glycerin.

### Procedure:

1. Put on your goggles and wear them throughout the lab. For each of the reactions, record your observations in the data table. Observations should include any evidence that a chemical reaction has taken place; the results of any test performed on any gases produced; the appearance of the products; and any other relevant data.
2. Obtain a 2 cm long piece of Mg ribbon. This length is approximate. The exact size is not important. Holding the Mg ribbon with the crucible tongs, carefully light the ribbon with the lab burner. Once lit hold the Mg ribbon over a watch glass to collect any ash and debris. When the Mg is done burning turn off the burner. **CAUTION:** Make sure hair is tied back. The tongs will be hot even after the Mg has finished burning. Do not touch them for at least five minutes. Place them on the watch glass to cool. Do not look directly at the Mg while it burns.

# Equation Writing And Predicting Products Answers

**Patrick E. McMahon, Rosemary  
McMahon, Bohdan Khomtchouk**

## Equation Writing And Predicting Products Answers:

**Survival Guide to General Chemistry** Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields

**General Chemistry Workbook** Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

*Jacaranda Chemistry 1 VCE Units 1 and 2, learnON and Print* Neale Taylor, Angela Stubbs, Robert Stokes, 2022-11-30 Developed by expert Victorian teachers for VCE students The NEW Jacaranda Chemistry VCE series continues to deliver curriculum aligned material that caters to students of all abilities Our expert author team of practising teachers and assessors ensures 100% coverage of the new VCE Chemistry Study Design 2023 2027

**Chemistry** Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02  
**Chemistry: The Central Science** Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Dalius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the

source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation *Laboratory Manual to Accompany Introductory Chemistry* Uno Kask, Otto W. Nitz, Martin G. Ondrus, Tammy J. Melton, 1993-07 This fifth edition of this laboratory manual emphasizes safety in the lab and discusses equipment requirements in the apparatus section at the beginning of each experiment It also features a revised art programme and explains the rationale for each experiment **Problem Exercises for General Chemistry** George Gilbert Long, Forrest C. Hentz, 1982 **Problem Exercises for General Chemistry** G. Gilbert Long, Forrest C. Hentz, 1986-08-14 A question answer supplement to accompany a 1 year general chemistry course This text is a multiple choice format and can be used with any standard general chemistry text The exercises emphasize the importance of problem solving and equation writing in the style used in general chemistry examinations and professional school aptitude examinations such as the MCAT and VCAT **Holt Chemistry** Ralph Thomas Myers, 2004 *General Chemistry* Raymond Chang, 2003 Publisher Description **Chemistry Problems and how to Solve Them** Paul Reheard Frey, 1969 **Introductory Chemistry** Nivaldo J. Tro, 2006 **General Chemistry, Reactions First** Kevin Revell, 2024-12-04 Revell's General Chemistry empowers students to grasp essential topics and concepts with more ease Using a friendly approach the text uses metaphors and relatable examples to demystify even the most challenging subjects in general chemistry *Foundations of College Chemistry* Morris Hein, 1970 *Chemistry* James E. Brady, Fred Senese, 2004-02-04 Publisher Description Modern University Chemistry, Custom Pub Norbert T. Porile, 1993 Basic Chemical Principles Edward I. Peters, 1988 **Textbook of Inorganic Chemistry** S. Young Tyree, Kerro Knox, 1961 *Contemporary Chemistry* Paul S. Cohen, Saul L. Geffner, 2004-01-10 To serve as a basal text for a high school chemistry course American Druggist, 1888

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