

Guided Practice: Equilibrium

Collision Theory & Reaction Rates

A - How Chemical Reactions Occur

- Using the collision model, explain at a molecular level why reactions proceed FASTER when concentrations increase.

B - Conditions that Affect Reaction Rates

- What is activation energy? After consulting your text (Fig 17.2 and 17.3) draw a graph showing a reaction pathway - labeling the activation energy within the graph.



- Use the collision theory to explain why reactions should occur more SLOWER at lower temperatures.
- How does a catalyst speed up a reaction? (Refer to the activation energy labeled in #2 in your answer.)
- Explain the difference between a homogeneous reaction and a heterogeneous reaction. Provide an example of each.
- List the factors that affect reaction rates; explain how each factor can affect the rate of a reaction:
 - a.
 - b.
 - c.
 - d.
- After reading the definition for equilibrium, provide an example of equilibrium in everyday life.
- Using your understanding of equilibrium, provide your own definition for chemical equilibrium.

Equilibrium Unit Webquest And Notes

ML Morrison



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