

Lab 5: RATE AND ORDER OF A CHEMICAL REACTION: IODINATION OF ACETONE

I. Introduction

The purpose of this experiment is to study a reaction that, in the vicinity of room temperature, proceeds at a moderate, relatively easily measured rate. For a given reaction, the rate typically increases with an increase in the concentration of any reactant. The relation between rate and concentration is a remarkably simple one in many cases. The objectives for this experiment include to determine the orders of reaction with respect to acetone and hydrogen ion; to confirm that the order of reaction with respect to iodine is zero; and to evaluate the rate constant k for the mixtures prepared at room temperature. For this experiment, we took safety precautions using laboratory materials and solutions, including wearing personal protective equipment.

II. Results Analysis

A. Data

Table 1. Reaction Rate Determination

Mixture Number	Initial Concentrations*			Time for reaction (s)	Temperature (°C)**	Reaction rate (M/s or mole/L · s)
	Acetone (M)	Iodine (M)	HCl (M)			
I (D.1)	5 mL	5 mL	5 mL	71.92	25	
II (D.2)	10 mL	5 mL	5 mL	35.10	25	
III (D.3)	5 mL	5 mL	10 mL	39.11	25	
IV (D.4)	5 mL	2.5 mL	5 mL	42.54	25	

Data Table 1. Group C - Iodination of Acetone: Time to Complete Reaction

Mixture	Temperature (°C)	Time (s) Took for Mixture Color to Disappear (seconds)
Mixture 1 (D.1)	25	71.92
Mixture 2 (D.2)	25	35.10
Mixture 3 (D.3)	25	39.11
Mixture 4 (D.4)	25	42.54

B. Discussion

Based on my results, I observed how the freezing point depression is directly related to the molecular weight of the solute in the solution. Using this relationship, the molecular weight of

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Experiments in General Chemistry Toby F. Block, 1986
Journal of the Chemical Society (Great Britain), 1959-10
Journal of the Chemical Society, 1951
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Microscale and Miniscale Organic Chemistry Laboratory Experiments Allen M. Schoffstall, Barbara A. Gaddis, Melvin L. Druelinger, 2000 This work offers a comprehensive introductory treatment of the organic laboratory techniques for handling glassware and equipment safety in the laboratory micro and mini scale experimental procedures theory of reactions and techniques applications and spectroscopy
The Industrial Arts Index, 1930
Journal of the American Chemical Society American Chemical Society, 1965 Proceedings of the Society are included in v 1 59 1879 1937
Annual Report United States. Public Health Service, 1940
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Journal of the Association of Official Agricultural Chemists Association of Official Analytical Chemists, 1925 Includes the Proceedings of the 30th 57th 1913 40 annual convention of the association Earlier proceedings were issued as Bulletins of the U S Dept of Agriculture Bureau of Chemistry
Bulletin of the Chemical Society of Japan Nihon Kagakkai, 1997
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Table of Contents Iodination Of Acetone Report

1. Understanding the eBook Iodination Of Acetone Report
 - The Rise of Digital Reading Iodination Of Acetone Report
 - Advantages of eBooks Over Traditional Books
2. Identifying Iodination Of Acetone Report
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Iodination Of Acetone Report
 - User-Friendly Interface
4. Exploring eBook Recommendations from Iodination Of Acetone Report
 - Personalized Recommendations
 - Iodination Of Acetone Report User Reviews and Ratings
 - Iodination Of Acetone Report and Bestseller Lists
5. Accessing Iodination Of Acetone Report Free and Paid eBooks
 - Iodination Of Acetone Report Public Domain eBooks
 - Iodination Of Acetone Report eBook Subscription Services
 - Iodination Of Acetone Report Budget-Friendly Options

6. Navigating Iodination Of Acetone Report eBook Formats
 - ePub, PDF, MOBI, and More
 - Iodination Of Acetone Report Compatibility with Devices
 - Iodination Of Acetone Report Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Iodination Of Acetone Report
 - Highlighting and Note-Taking Iodination Of Acetone Report
 - Interactive Elements Iodination Of Acetone Report
8. Staying Engaged with Iodination Of Acetone Report
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Iodination Of Acetone Report
9. Balancing eBooks and Physical Books Iodination Of Acetone Report
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Iodination Of Acetone Report
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Iodination Of Acetone Report
 - Setting Reading Goals Iodination Of Acetone Report
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Iodination Of Acetone Report
 - Fact-Checking eBook Content of Iodination Of Acetone Report
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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