

CHEMICAL BONDS

Ionic Bonding

Metals and non-metals react together to form _____ compounds. They are held together by ionic bonds. Ionic bonding is a force of _____ attraction between two _____ atoms (called _____). An atom becomes charged by either donating or accepting _____ electrons from another atom. A valence electron is one from the _____ shell of the atom.

Since opposites attract, there is an attraction between the positively charged and negatively charged ions that holds them together.



If an atom _____ an electron, it becomes positively charged. This is then called a _____. _____ always donate electrons. If an atom accepts an electron, it becomes _____. This is called an _____. Non-metals always accept electrons.

Atoms do this in order to become '_____'. A stable atom is one which has obtained a _____ of electrons.

How do we know which atoms donate and which accept electrons?

Sodium has _____ lone electron in the outer shell. It is easiest for this atom to _____ one electron than to try and find seven more to fill the outer shell.



All _____ donate electrons and therefore become cations.

All _____ accept electrons and therefore become anions.

Chlorine has _____ electrons in its outer shell. Therefore, it is easiest for Chlorine to _____ one electron than to _____ to try and donate seven.



Ionic And Covalent Bonds Study Guide

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Table of Contents Ionic And Covalent Bonds Study Guide

1. Understanding the eBook Ionic And Covalent Bonds Study Guide
 - The Rise of Digital Reading Ionic And Covalent Bonds Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic And Covalent Bonds Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic And Covalent Bonds Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic And Covalent Bonds Study Guide
 - Personalized Recommendations
 - Ionic And Covalent Bonds Study Guide User Reviews and Ratings
 - Ionic And Covalent Bonds Study Guide and Bestseller Lists
5. Accessing Ionic And Covalent Bonds Study Guide Free and Paid eBooks
 - Ionic And Covalent Bonds Study Guide Public Domain eBooks
 - Ionic And Covalent Bonds Study Guide eBook Subscription Services
 - Ionic And Covalent Bonds Study Guide Budget-Friendly Options
6. Navigating Ionic And Covalent Bonds Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Ionic And Covalent Bonds Study Guide Compatibility with Devices
 - Ionic And Covalent Bonds Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic And Covalent Bonds Study Guide
 - Highlighting and Note-Taking Ionic And Covalent Bonds Study Guide
 - Interactive Elements Ionic And Covalent Bonds Study Guide
8. Staying Engaged with Ionic And Covalent Bonds Study Guide

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic And Covalent Bonds Study Guide
9. Balancing eBooks and Physical Books Ionic And Covalent Bonds Study Guide
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic And Covalent Bonds Study Guide
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ionic And Covalent Bonds Study Guide
- Setting Reading Goals Ionic And Covalent Bonds Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ionic And Covalent Bonds Study Guide
- Fact-Checking eBook Content of Ionic And Covalent Bonds Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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