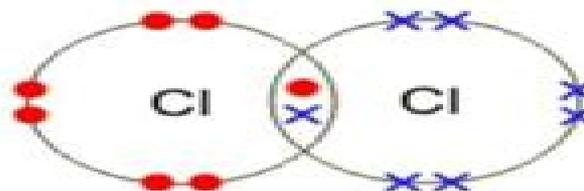


### Covalent bonding

A non-metal metal and another non-metal can bond together by **sharing** electrons so that both electrons can achieve a **full outer shell**. This is called **covalent bonding**. To represent the atoms involved in the bonding we draw crosses for the electrons on one of the atoms and dots for the electrons on the other atoms as shown here.



### Instructions

**NB:** It may be easiest to draw each atom out individually first to work out how many electrons will be shared

1. Draw the circles (outer shells) and overlap them (like a Venn diagram)
2. Write the symbols of the atoms involved in the middle of each circle
3. Draw the **electrons** for one atom as dots (o) and electrons for the other atom as crosses (x).
4. Any shared electrons are drawn in the **overlapping** section.
5. **Count up the electrons for EACH atom and check that they each now have a full outer shell** (2 electrons for the first shell and 8 for the next 2 shells)

**TASK:** Draw dot and cross diagrams for the following covalent compounds.

$F_2$		$CH_4$	
$HCl$		$CO_2$	
$H_2S$		$H_2O$	
$HF$		$NH_3$	

**Challenge** -  $O_2$

# Ionic And Metallic Bonding Questions

**William D. Callister, Jr., David G.  
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